

5

ELECTRONS IN ATOMS

SECTION 5.1 MODELS OF THE ATOM (pages 127–132)

This section summarizes the development of atomic theory. It also explains the significance of quantized energies of electrons as they relate to the quantum mechanical model of the atom.

► The Development of Atomic Models (pages 127–128)

1. Complete the table about atomic models and the scientists who developed them.

Scientist	Model of Atom
Dalton	The atom is a solid indivisible mass.
Thomson	The atom is a ball of positive charge with electrons stuck into the ball.
Rutherford	Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space.
Bohr	Electrons are arranged in concentric circular paths around the nucleus.

2. Is the following sentence true or false? The electrons in an atom can exist between energy levels. false

► The Bohr Model (pages 128–129)

3. What is a small, discrete unit of energy called?
It is called a quantum.
4. Circle the letter of the term that completes the sentence correctly: A quantum of energy is the amount of energy required to
- move an electron from its present energy level to the next lower one
 - maintain an electron in its present energy level
 - c.** move an electron from its present energy level to the next higher one
5. In general, the higher the electron is on the energy ladder, the farther it is from the nucleus.

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Table of Contents Electron Arrangement In Atoms Guided Reading Answers

1. Understanding the eBook Electron Arrangement In Atoms Guided Reading Answers
 - The Rise of Digital Reading Electron Arrangement In Atoms Guided Reading Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Arrangement In Atoms Guided Reading Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Arrangement In Atoms Guided Reading Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Arrangement In Atoms Guided Reading Answers

- Personalized Recommendations
 - Electron Arrangement In Atoms Guided Reading Answers User Reviews and Ratings
 - Electron Arrangement In Atoms Guided Reading Answers and Bestseller Lists
5. Accessing Electron Arrangement In Atoms Guided Reading Answers Free and Paid eBooks
 - Electron Arrangement In Atoms Guided Reading Answers Public Domain eBooks
 - Electron Arrangement In Atoms Guided Reading Answers eBook Subscription Services
 - Electron Arrangement In Atoms Guided Reading Answers Budget-Friendly Options
 6. Navigating Electron Arrangement In Atoms Guided Reading Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Arrangement In Atoms Guided Reading Answers Compatibility with Devices
 - Electron Arrangement In Atoms Guided Reading Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Arrangement In Atoms Guided Reading Answers
 - Highlighting and Note-Taking Electron Arrangement In Atoms Guided Reading Answers
 - Interactive Elements Electron Arrangement In Atoms Guided Reading Answers
 8. Staying Engaged with Electron Arrangement In Atoms Guided Reading Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Arrangement In Atoms Guided Reading Answers
 9. Balancing eBooks and Physical Books Electron Arrangement In Atoms Guided Reading Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Arrangement In Atoms Guided Reading Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Electron Arrangement In Atoms Guided Reading Answers
 - Setting Reading Goals Electron Arrangement In Atoms Guided Reading Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Electron Arrangement In Atoms Guided Reading Answers

- Fact-Checking eBook Content of Electron Arrangement In Atoms Guided Reading Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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