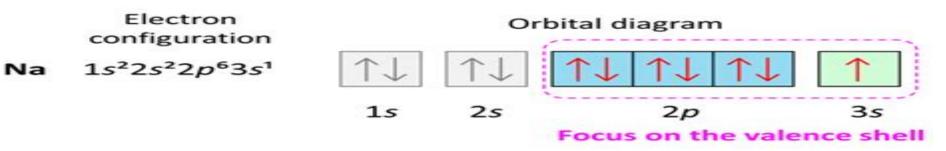
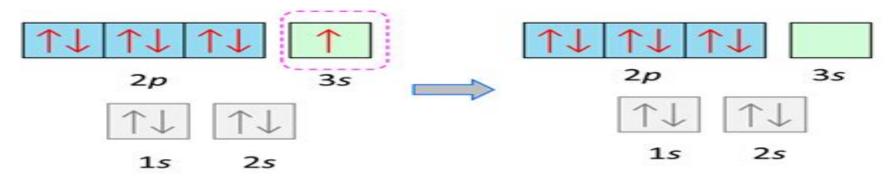
Electron Configuration of Main Group Cations



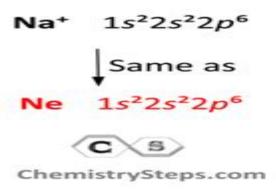
The charge is +1, therefore, one electron is lost:

The electron must be from the outermost valence orbital $-3s^1$



Na
$$1s^2 2s^2 2p^6 3s^1$$

Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are *isoelectronic* with the previous noble gas.



Electron Configurations Of Atoms And Ion Lab

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