

Name: \_\_\_\_\_ Date: \_\_\_\_\_

## Electron configuration worksheet with answer key

Write the full electron configurations for each element.

- Potassium (K)  $1s^1 2s^2 2p^6 3s^2 3p^6 4s^1$
- Sulfur (S)  $1s^2 2s^2 2p^6 3s^2 3p^4$
- Fluorine (F)  $1s^2 2s^2 2p^5$
- Silicon (Si)  $1s^2 2s^2 2p^6 3s^2 3p^2$

Write the shorthand (noble gas) electron configurations.

- Calcium (Ca)  $[\text{Ar}] 4s^2$
- Iron (Fe)  $[\text{Ar}] 4s^2 3d^6$
- Nickel (Ni)  $[\text{Ar}] 4s^2 3d^8$
- Copper (Cu)  $[\text{Ar}] 4s^1 3d^{10}$

Identify elements based on their electron configurations.

- $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$  K (Potassium)
- $[\text{Ne}] 3s^2 3p^3$  P (Phosphorus)
- $[\text{Ar}] 4s^2 3d^8$  (Nickel)
- $[\text{Kr}] 5s^1$  (Rubidium)

# Electron Configuration Lab Answers

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