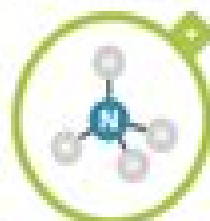


# POLYATOMIC IONS: NAMES, FORMULAE & CHARGES

A polyatomic ion is a charged species consisting of two or more atoms covalently bonded together. Here's a guide to some of the most common examples!



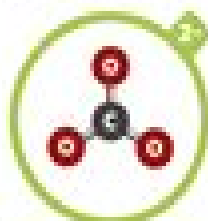
**AMMONIUM**

Formula:  $\text{NH}_4^+$



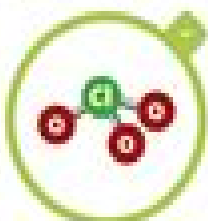
**ACETATE**

Formula:  $\text{C}_2\text{H}_3\text{O}_2^-$



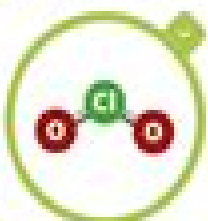
**CARBONATE**

Formula:  $\text{CO}_3^{2-}$



**CHLORATE**

Formula:  $\text{ClO}_3^-$



**CHLORITE**

Formula:  $\text{ClO}_2^-$



**CHROMATE**

Formula:  $\text{CrO}_4^{2-}$



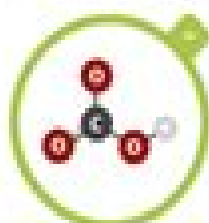
**CYANIDE**

Formula:  $\text{CN}^-$



**DICHROMATE**

Formula:  $\text{Cr}_2\text{O}_7^{2-}$



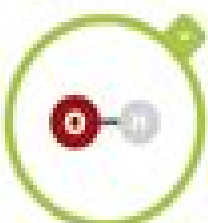
**HYDROGEN CARBONATE**

Formula:  $\text{HCO}_3^-$



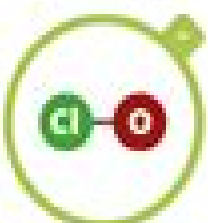
**HYDROGEN SULFATE**

Formula:  $\text{HSO}_4^-$



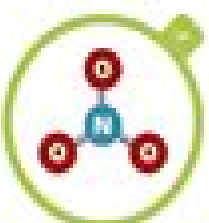
**HYDROXIDE**

Formula:  $\text{OH}^-$



**HYPOCHLORITE**

Formula:  $\text{ClO}^-$



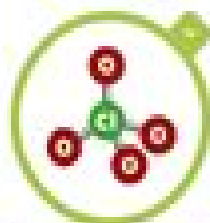
**NITRATE**

Formula:  $\text{NO}_3^-$



**NITRITE**

Formula:  $\text{NO}_2^-$



**PERCHLORATE**

Formula:  $\text{ClO}_4^-$



**PERMANGANATE**

Formula:  $\text{MnO}_4^-$



**PEROXIDE**

Formula:  $\text{O}_2^{2-}$



**PHOSPHATE**

Formula:  $\text{PO}_4^{3-}$



**SULFATE**

Formula:  $\text{SO}_4^{2-}$



**SULFITE**

Formula:  $\text{SO}_3^{2-}$



**THIOSULFATE**

Formula:  $\text{S}_2\text{O}_3^{2-}$



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# Guide To Polyatomic Ions

**Thirumalaisamy P.  
Velavancorresponding**



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