Explain how you identified the limiting and excess reactant based only on your qualitative observations (not quantitative measurements or calculations). Type your answer into the box. Use complete sentences.

According to observation the initial color of browningsgreegy purple is yellow in color after adding the NaOH, it turns violet. After Changing the color yellow from purple indicates acid is completely consumed so excess of base addition changes the color into violet.

Table C. Calculations to Determine Theoretical Yields

Description	Calculated Value
Volume of NaOH solution	22.20 mL
Initial moles of NaOH	0.002451 moles
Initial moles of H ₃ PO ₄	0.002602 moles
Mass of H ₂ O produced if NaOH runs out first	0.04414 g
Mass of H ₂ O produced if H ₃ PO ₄ runs out first	1.265 g
Theoretical yield of H ₂ O (in grams)	1.265 g
Theoretical yield of Na ₃ PO ₄ (in grams)	0.4267 g

Table D. Identification of Limiting and Excess Reactants from Calculations

Description	ption Reactant Name or Formula	
Limiting Reactant	Phosphoric Acid	
Excess Reactant	Sodium Hydroxide	

Explain how you identified the limiting and excess reactant based on your calculations. Be specific. Type your answer into the box. Use complete sentences.

According to our calculation the reacting equivalent of H₃PO₄ is less than NaOH. Therefore H₂PO₄ is limiting reactant so product formation depends on H₃PO₄ because phosphoric acid is completely consumed.

Formal Lab Report On Limiting Reactant

Donald J. Wink, Sharon Fetzer-Gislason, Julie Ellefson Kuehn Formal Lab Report On Limiting Reactant:

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