

1. According to the Bronsted—Lowry definition, which chemical species can function *both* as an acid and as a base?

- $\text{Cl}^-$
- $\text{SO}_4^{2-}$
- $\text{NH}_4^+$
- $\text{HCO}_3^-$
- $\text{H}_3\text{O}^+$

2. In the reaction  $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$  which is an acid—base conjugate pair?

- $\text{H}_2\text{O}$  and  $\text{HCN}$
- $\text{H}_2\text{O}$  and  $\text{OH}^-$
- $\text{CN}^-$  and  $\text{H}_2\text{O}$
- $\text{HCN}$  and  $\text{OH}^-$

3. Given that  $\text{HX}$  is a stronger Bronsted acid than  $\text{HY}$  in aqueous solution, which is true of a 1 M solution of  $\text{NaX}$ ?

- It is less basic than a 1 M solution of  $\text{NaY}$ .
- It is more basic than a 1 M solution of  $\text{NaY}$ .
- It yields a neutral solution.
- It is more concentrated than a 1 M solution of  $\text{NaY}$ .

4. The pH of a 0.03 M  $\text{HCl}$  solution is

- 1.5
- 2.5
- 3.5
- 12.5

5. The pH of a  $1.0 \times 10^{-2}$  M  $\text{Ba}(\text{OH})_2$  solution at  $25^\circ\text{C}$  is

- 2.7
- 3.0
- 11.0
- 11.3

6. Which is the strongest acid?

- $\text{HClO}$
- $\text{HClO}_2$
- $\text{HClO}_3$
- $\text{HClO}_4$

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

- $\text{NaC}_2\text{H}_3\text{O}_2$
- $\text{NaNO}_3$
- $\text{NH}_4\text{Cl}$
- $\text{BaSO}_4$

8. The *weakest* of the bases listed is

Acid	Conjugate Base	$K_a$ (Ionization Constant of Acid)
$\text{HCl}$	$\text{Cl}^-$	100% ionized
$\text{HSO}_4^-$	$\text{SO}_4^{2-}$	$1.2 \times 10^{-2}$
$\text{H}_2\text{S}$	$\text{HS}^-$	$5.7 \times 10^{-8}$
$\text{HS}^-$	$\text{S}^{2-}$	$1.2 \times 10^{-13}$

- $\text{Cl}^-$
- $\text{CN}^-$
- $\text{HS}^-$
- $\text{S}^{2-}$
- $\text{SO}_4^{2-}$

9. The oxide of which element will react with water to form the strongest acid?

	Main Groups							
Period	I	II	III	IV	V	VI	VII	O
First								
Second	X	Y	Z	P		Q	S	U
Third	W					R	T	M

- W
- M
- P
- R
- Z

10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate  $\text{KC}_2\text{H}_3\text{O}_2$  is less alkaline than a 0.10 M solution of potassium cyanide,  $\text{KCN}$ ?

- Hydrocyanic acid is a weaker acid than acetic acid.
- Hydrocyanic acid is less soluble in water than acetic acid.
- Cyanides are less soluble than acetates.
- Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M  $\text{NaOH}$ , the properties of the solution at the equivalence point will correspond exactly to the properties of

- a 0.100 M sodium solution.
- a 0.0500 M sodium hydroxide solution.
- a 0.0500 M benzoic acid solution.
- a 0.0500 M sodium benzoate solution.

# Inorganic Chemistry 2009 Acs Exam Study Guide

**Michael Seilmaier**



## **Inorganic Chemistry 2009 Acs Exam Study Guide:**

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