

Assignment #2 – Compound Names and Formulas
Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

- copper (II) oxide CuO
- mercury (I) oxide Hg₂O
- gold (III) chloride AuCl₃
- thallium (III) bromide TlBr₃
- bismuth (V) oxide Bi₂O₅
- terbium (III) oxide Tb₂O₃
- uranium (VI) oxide UO₃
- protactinium (IV) bromide PaBr₄
- cerium (III) oxide Ce₂O₃
- arsenic (V) sulphide As₂S₅

- manganese (III) oxide Mn₂O₃
- vanadium (II) bromide VBr₂
- niobium(V) oxide Nb₂O₅
- titanium (III) oxide Ti₂O₃
- titanium (III) nitride TiN
- iron (II) oxide FeO
- cobalt (II) phosphide Co₃P₂
- tin (II) oxide SnO
- thulium (II) bromide TmBr₂
- copper (II) bromide CuBr₂

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

- SnCl₄ Tin (IV) Chloride
- BiBr₅ Bismuth (V) Bromide
- PoO₂ Polonium (IV) Oxide
- PbI₂ Lead (II) Iodide
- HgO Mercury (II) Oxide
- HgCl Mercury (I) Chloride
- Au₂O₃ Gold (III) Oxide
- FeCl₂ Iron (II) Chloride

- PdF₄ Palladium (IV) Fluoride
- Os₂O₃ Osmium (III) Oxide
- MoBr₂ Molybdenum (II) Bromide
- VCl₅ Vanadium (V) Chloride
- Mn₂O₃ Manganese (III) Oxide
- CoO Cobalt (II) Oxide
- Np₂O₃ Neptunium (III) Oxide
- V₂O₅ Vanadium (V) Oxide

Formula Of Ionic Compound Lab Answer Key

B Lingard



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Table of Contents Formula Of Ionic Compound Lab Answer Key

1. Understanding the eBook Formula Of Ionic Compound Lab Answer Key
 - The Rise of Digital Reading Formula Of Ionic Compound Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Formula Of Ionic Compound Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Formula Of Ionic Compound Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Formula Of Ionic Compound Lab Answer Key
 - Personalized Recommendations
 - Formula Of Ionic Compound Lab Answer Key User Reviews and Ratings
 - Formula Of Ionic Compound Lab Answer Key and Bestseller Lists

5. Accessing Formula Of Ionic Compound Lab Answer Key Free and Paid eBooks
 - Formula Of Ionic Compound Lab Answer Key Public Domain eBooks
 - Formula Of Ionic Compound Lab Answer Key eBook Subscription Services
 - Formula Of Ionic Compound Lab Answer Key Budget-Friendly Options
6. Navigating Formula Of Ionic Compound Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Formula Of Ionic Compound Lab Answer Key Compatibility with Devices
 - Formula Of Ionic Compound Lab Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Formula Of Ionic Compound Lab Answer Key
 - Highlighting and Note-Taking Formula Of Ionic Compound Lab Answer Key
 - Interactive Elements Formula Of Ionic Compound Lab Answer Key
8. Staying Engaged with Formula Of Ionic Compound Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Formula Of Ionic Compound Lab Answer Key
9. Balancing eBooks and Physical Books Formula Of Ionic Compound Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Formula Of Ionic Compound Lab Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Formula Of Ionic Compound Lab Answer Key
 - Setting Reading Goals Formula Of Ionic Compound Lab Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Formula Of Ionic Compound Lab Answer Key
 - Fact-Checking eBook Content of Formula Of Ionic Compound Lab Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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