

Example:

25. Three grams of carbon and 12 g of oxygen are present and volume 200 g of CO_2 are obtained with no further reaction of carbon. Find

$$25 \text{ g } \text{C} \times \frac{1 \text{ mol C}}{12 \text{ g C}} \times \frac{8 \text{ mol CO}_2}{15 \text{ mol C}} \times \frac{44 \text{ g CO}_2}{1 \text{ mol CO}_2} = \boxed{74.67 \text{ g CO}_2}$$

26. When 15 g of Al reacts with the oxygen of 25 g of O_2 , 20.5 g of Al_2O_3 are produced. What is the percentage yield of this reaction?



$$15 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times \frac{102 \text{ g Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} = 50.67 \text{ g Al}_2\text{O}_3 \text{ (Theoretical Yield)}$$

$$\% \text{ Yield} = \frac{20.5 \text{ g Al}_2\text{O}_3}{50.67 \text{ g Al}_2\text{O}_3} \times 100 = \boxed{73.8\%}$$

27. Following to problem 26, what is the limiting reagent when 200 g of Al and 25 g of O_2 react with 20.5 g of Al_2O_3 ?

$$200 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{3 \text{ mol O}_2}{4 \text{ mol Al}} \times \frac{32 \text{ g O}_2}{1 \text{ mol O}_2} = 3.0 \text{ g O}_2$$

$$25 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{2 \text{ mol Al}}{3 \text{ mol O}_2} \times \frac{27 \text{ g Al}}{1 \text{ mol Al}} = \boxed{2.78 \text{ g Al}} \text{ (maximum product)}$$

Limiting
Reagent = Al

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Katrin Zwirgmaier

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