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<b>Pearson Edexcel Level 3 GCE</b>							
<b>Thursday 25 May 2023</b>							
Afternoon (Time: 1 hour 30 minutes)				Paper reference		<b>9FM0/01</b>	
<b>Further Mathematics</b> <b>Advanced</b> <b>PAPER 1: Core Pure Mathematics 1</b>							
<b>You must have:</b> Mathematical Formulae and Statistical Tables (Green), calculator						Total Marks	

Candidates may use any calculator permitted by Pearson regulations. Calculators must not have the facility for algebraic manipulation, differentiation and integration, or have retrievable mathematical formulae stored in them.

#### Instructions

- Use **black** ink or ball-point pen.
- If pencil is used for diagrams/sketches/graphs it must be dark (HB or B).
- **Fill in the boxes** at the top of this page with your name, centre number and candidate number.
- Answer **all** questions and ensure that your answers to parts of questions are clearly labelled.
- Answer the questions in the spaces provided  
– *there may be more space than you need.*
- You should show sufficient working to make your methods clear. Answers without working may not gain full credit.
- Inexact answers should be given to three significant figures unless otherwise stated.

#### Information

- A booklet 'Mathematical Formulae and Statistical Tables' is provided.
- There are 8 questions in this question paper. The total mark for this paper is 75.
- The marks for **each** question are shown in brackets  
– *use this as a guide as to how much time to spend on each question.*

#### Advice

- Read each question carefully before you start to answer it.
- Try to answer every question.
- Check your answers if you have time at the end.

Turn over ➤



# May 2015 Edexcel Style Maths Paper 1 Mark Scheme

**Iseb Iseb**



**May 2015 Edexcel Style Maths Paper 1 Mark Scheme:**

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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients  $\text{Cu} : \text{MnO}_4^{-}$  is 4 : 1

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web 1 practice problem balance the redox reaction in basic solution  $\text{NO}_2^{-} + \text{Al}(\text{OH})_3 + \text{OH}^{-} \rightarrow \text{AlO}_2^{-} + \text{NH}_3 + \text{H}_2\text{O}$  practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of  $\text{OH}^{-}$  before the reaction is balanced  
topic next topic

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web google classroom for the following redox equation  $\text{MnO}_4^{-} + \text{SO}_2 + 2\text{H}^{+} \rightarrow \text{Mn}^{2+} + \text{HSO}_4^{-} + \text{H}_2\text{O}$  what will be the stoichiometric coefficient of  $\text{HSO}_4^{-}$  if the above equation is balanced in an acidic medium stuck use a hint report a problem



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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN} + 2\text{e}^-$   $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{H}_2\text{O} \rightarrow 2\text{AuCN} + 2\text{OH}^-$  multiplied by a factor of 2  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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stages first bromine reacts with  $\text{OH}^-$  ions to form bromide ions and  $\text{BrO}^-$  ions second the  $\text{BrO}^-$  ions react to form bromide ions and bromate  $\text{V}^-$  ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{H}^+ \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+} + \text{H}_2\text{O}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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