

15 Hydrocarbons

Alkanes

Chemical Reactivity

Alkanes do not react with many reagents.

This is because the C-C bond and the C-H bond are relatively strong and so cannot form stronger bonds with other elements that easily and there are no polar side groups.

Their main reactions are combustion reactions in air although they can react with halogens in the presence of UV light.

Combustion

Complete Combustion

In excess oxygen alkanes will burn with complete combustion.

The products of complete combustion are CO_2 and H_2O .



Alkanes readily burn in the presence of oxygen. This combustion of alkanes is highly exothermic, explaining their use as **fuels**.

Reaction of alkanes with bromine / chlorine in UV light

Synthesis of chloroalkanes

In the presence of **UV light** alkanes react with chlorine to form a mixture of products with the halogens substituting hydrogen atoms.

Overall Reaction



This is the overall reaction, but a more complex mixture of products is actually formed.

*To understand this reaction fully we must look in detail at how it proceeds step by step. This is called its **mechanism**.*

The **MECHANISM** for this reaction is called a **FREE RADICAL SUBSTITUTION**.

It proceeds via a series of steps:
STEP ONE: Initiation
STEP TWO: Propagation
STEP THREE: Termination

STEP ONE Initiation

Essential condition: UV light



The UV light supplies the energy to break the Cl-Cl bond. It is broken in preference to the others as it is the weakest.

The bond has broken in a process called **homolytic fission**.

*each atom gets **one** electron from the covalent bond*

When a bond breaks by homolytic fission it forms **Free Radicals**.

Free Radicals do not have a charge and are represented by a



DEFINITION

A **Free Radical** is a reactive species which possess an **unpaired electron**.

STEP TWO Propagation



The chlorine free radicals are very reactive and remove an H from the methane leaving a methyl free radical

The methyl free radical reacts with a Cl_2 molecule to produce the main product and another Cl free radical

All propagation steps have a **free radical** in the **reactants** and in the **products**.

As the Cl free radical is regenerated, it can react with several more alkane molecules in a **CHAIN REACTION**.

Hydrocarbon Compounds Study Guide Answers

Lauren Gardner



Hydrocarbon Compounds Study Guide Answers:

This book delves into Hydrocarbon Compounds Study Guide Answers. Hydrocarbon Compounds Study Guide Answers is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Hydrocarbon Compounds Study Guide Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Hydrocarbon Compounds Study Guide Answers
- Chapter 2: Essential Elements of Hydrocarbon Compounds Study Guide Answers
- Chapter 3: Hydrocarbon Compounds Study Guide Answers in Everyday Life
- Chapter 4: Hydrocarbon Compounds Study Guide Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Hydrocarbon Compounds Study Guide Answers. This chapter will explore what Hydrocarbon Compounds Study Guide Answers is, why Hydrocarbon Compounds Study Guide Answers is vital, and how to effectively learn about Hydrocarbon Compounds Study Guide Answers.
3. In chapter 2, the author will delve into the foundational concepts of Hydrocarbon Compounds Study Guide Answers. This chapter will elucidate the essential principles that must be understood to grasp Hydrocarbon Compounds Study Guide Answers in its entirety.
4. In chapter 3, this book will examine the practical applications of Hydrocarbon Compounds Study Guide Answers in daily life. The third chapter will showcase real-world examples of how Hydrocarbon Compounds Study Guide Answers can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Hydrocarbon Compounds Study Guide Answers in specific contexts. This chapter will explore how Hydrocarbon Compounds Study Guide Answers is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Hydrocarbon Compounds Study Guide Answers. This chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Hydrocarbon Compounds Study Guide Answers.

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