

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.



- **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.
 - **Entropy** is a measure of the disorder or randomness of a system.

$$\Delta S = \frac{Q}{T}$$

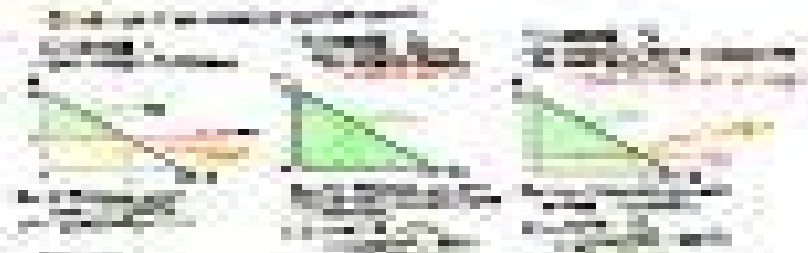
 where ΔS is the change in entropy, Q is the heat added, and T is the temperature.



- **Third Law of Thermodynamics**: The entropy of a perfect crystal at absolute zero is zero.
 - **Heat Engines** convert heat into work. The efficiency of a heat engine is given by:

$$\eta = \frac{W}{Q_{in}}$$

- **Heat Pumps and Refrigerators** transfer heat from a cold reservoir to a hot reservoir.
 - **Carnot Cycle** is the most efficient cycle possible between two temperatures.



- **Statistical Mechanics** connects the microscopic properties of particles to the macroscopic properties of a system.
 - **Boltzmann's Equation** relates the average energy of a particle to the temperature:

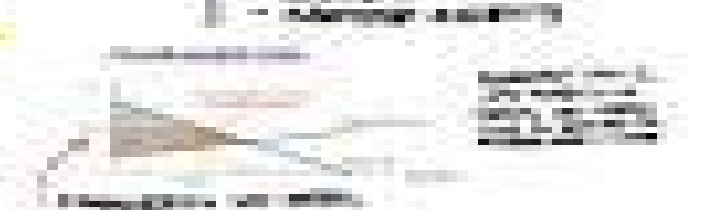
$$\langle E \rangle = \frac{3}{2} k_B T$$

- **Maxwell's Equations** describe the relationship between electric and magnetic fields.
 - **Electromagnetic Induction** is the process by which a changing magnetic field induces an electric current.

- **Quantum Mechanics** describes the behavior of particles at the atomic and subatomic level.
 - **Schrodinger's Equation** is the fundamental equation of quantum mechanics.

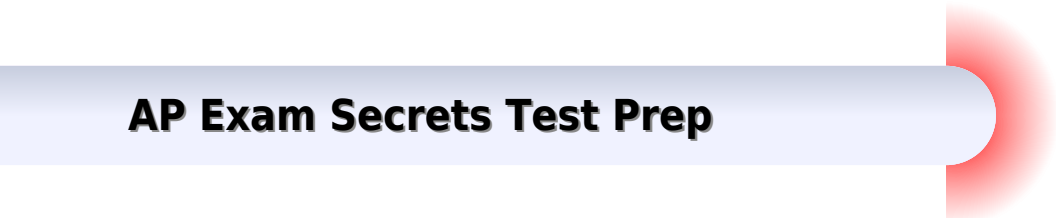


- **Relativity** describes the relationship between space and time.
 - **Einstein's Theory of Relativity** shows that the laws of physics are the same for all observers.



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