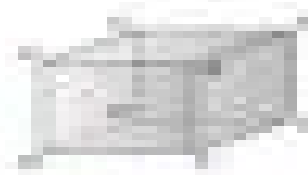


Answer following



- 1. Substrate and temperature: 20°C 30°C 40°C
- 2. A graph of initial velocity: 100 mg/min 200 mg/min 300 mg/min
- 3. A graph of initial velocity: 100 mg/min 200 mg/min 300 mg/min

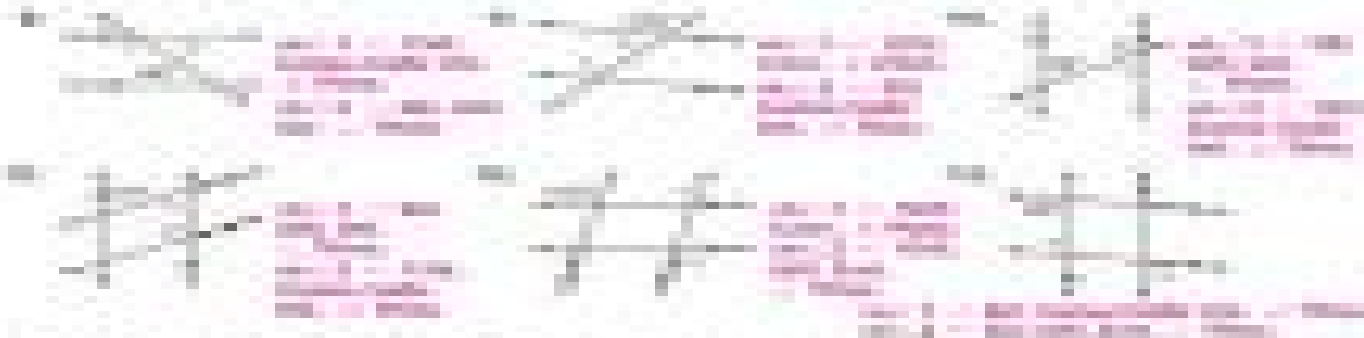
Calculate all values of K_m and V_{max} for the following reaction. Assume Michaelis-Menten kinetics and that the reaction is reversible.

- 1. $K_m = 100\text{ mg/min}$ $V_{max} = 300\text{ mg/min}$
- 2. $K_m = 100\text{ mg/min}$ $V_{max} = 300\text{ mg/min}$
- 3. $K_m = 100\text{ mg/min}$ $V_{max} = 300\text{ mg/min}$
- 4. $K_m = 100\text{ mg/min}$ $V_{max} = 300\text{ mg/min}$



Answer 100

Graph the following data. Assume the Michaelis-Menten kinetics and that the reaction is reversible.



Graph the following data. Assume the Michaelis-Menten kinetics and that the reaction is reversible.



Michaelis-Menten kinetics: $V = \frac{V_{max}[S]}{K_m + [S]}$

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Jan Fair



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