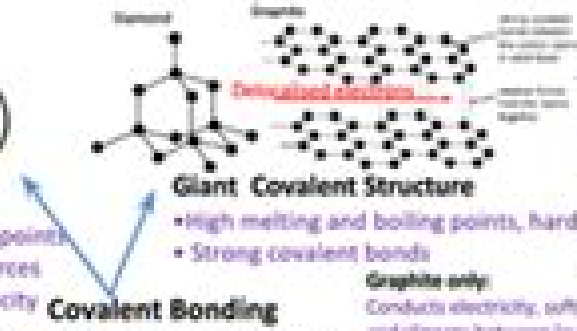
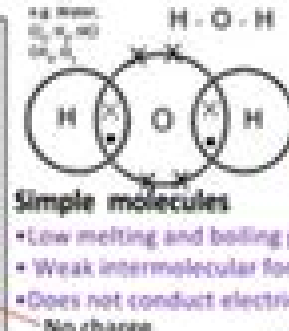




Periodic Table of the Elements

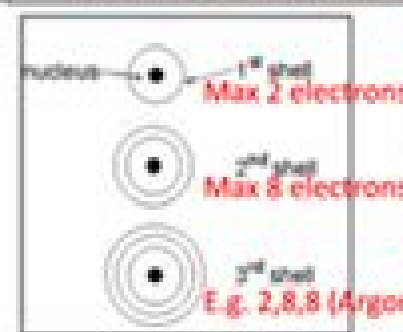
- metals
- non-metals
- metalloids
- transition metals
- alkali metals
- alkali earth metals
- transition metals
- lanthanoids
- actinoids

An atom:
 Number of Protons = Number of Electrons
 Mass number = Protons + Neutrons

C2 Chemistry

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
 E.g. Lithium 2,1 is a group 1 element.

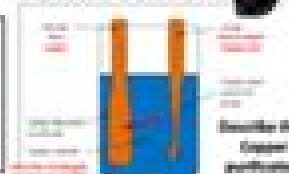


Protons and neutrons both have a mass of 1

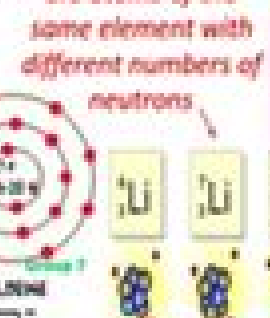
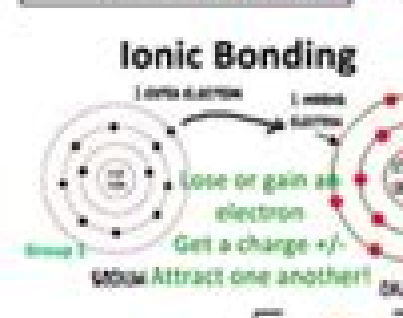
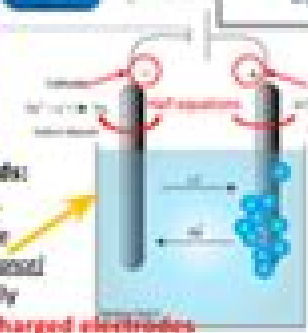
ISOTOPES are atoms of the same element with different numbers of neutrons

OIL RIG

Oxidation is loss (of electrons)
 Reduction is gain (of electrons)



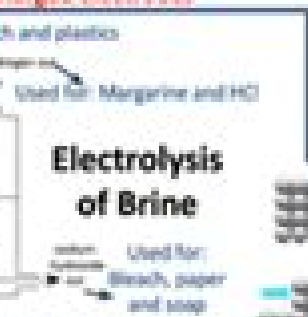
Electrolysis:
 Splitting up a substance using electricity



When there are 2 possible products The less reactive forms at the electrode

Electrolysis needs:

- 2 electrodes
- An electrolyte (ionic or molten substance)
- A power supply

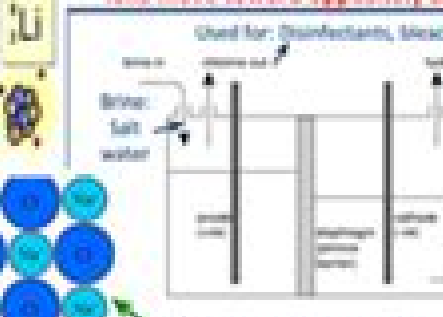


Metal ions only arranged in layers

ELECTROSTATIC ATTRACTION between +IONS and ELECTRONS

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well



NANO-TECHNOLOGY
 Means REALLY REALLY REALLY small!!!!

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Marco Cascella



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