

EXPERIMENT: INVESTIGATING THE DENSITY OF WATER AND ICE

Syllabus reference 8.4.1

INTRODUCTION

There are a number of important properties that assist in identifying elements and pure compounds. These include density, melting point and boiling point.

The density of a substance is the measure of its compactness and is calculated from the ratio:

$$\text{density} = \frac{\text{mass (g)}}{\text{volume (mL)}}$$

In this experiment you will:

- a measure the mass of several different volumes of water and use these to calculate the density of water
- b use a similar procedure to calculate the density of ice.

EQUIPMENT

- 100 mL beaker
- wide diameter measuring cylinder
- electronic balance
- distilled water
- ice cubes (made with distilled water and small enough to fit into the measuring cylinder)
- 10 mL pipette
- forceps or glass rod
- thermometer

Part A: Density of water

PROCEDURE

- 1 Accurately measure the mass of the empty beaker. Record this in the table below.
- 2 Add a known volume of distilled water (use the pipette) to the beaker and reweigh. Record this in the table.
- 3 Measure the temperature of the water.
- 4 Repeat this procedure until you have at least 5 measurements (use 10 mL increments from the starting volume). Record all measurements in the table.

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