Chemistry of Life

## Study Guide B

#### Answer Key

## SECTION 1. ATOMS, IONS, AND MOLECULES

- I. An element is a certain type of atom.
- Sketch should resemble one of the illustrations in Figure 1.2 in Section 1.
   Nucleus should be in the center with protons and neutrons. Electrons should be labeled in a ring surrounding the nucleus.
- A compound is composed of atoms of different elements.
- an atom that has gained or lost one or more electrons
- the number of protons no longer equals the number of electrons
- Sketch should resemble the first illustration in Figure 1.3, with one atom losing an electron (becoming a positive ion) and the other accepting that electron (becoming a negative ion).
- A bond that forms between two atoms that share a pair of electrons.
- the number of electron pairs that it needs to share to full its outer energy level
- 9. molecule
- 10. compound
- 11. element
- 1.2. Sept.
- Ionic bonds form between oppositelycharged ions; covalent bonds form when atoms share a pair of electrons.

## SECTION 2. PROPERTIES OF WATER

- a molecule with a slightly positively charged region and a slightly negatively charged region
- The oxygen nucleus pulls the electrons in the molecule more strongly than the hydrogen atoms. The hydrogen atoms have a slight positive charge and the oxygen atom has a slight negative charge.
- an attraction between a slightly positive hydrogen and a slightly negative atom
- between the slightly positive hydrogen atom of one water molecule and the

slightly negative oxygen atom of another water molecule

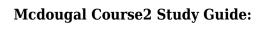
- A large amount of energy is needed to produce an increase in water temperature; water resists changes in temperature.
- 6. Water molecules "stick" to each other.
- Water molecules "stick" to other substances.
- A solvent is present in greater concentration and dissolves the solute.
- 9. ions and polar molecules
- 10. n-onpolar molecules, such as fats and oils
- Effect on H\* concentration: acids increase H\* concentration; bases decrease H\* concentration
   Effect on pH: acids lower pH; bases raise
- 12. Sketch should resemble Visual Vocab in Section 2 of the text. Solution: homogeneous mixture of substances; solvent: substance present in greater amount in which other substances dissolve; solute: substance that dissolves in a solvent.

## SECTION 3. CARBON-BASED MOLECULES

- Carbon atoms are the basis of the molecules that make up most living things.
- Carbon atoms can form covalent bonds with up to four other atoms, including other carbon atoms.
- Sketches should resemble Figure 3.1 in Section 3.
- broken down as a source of chemical energy; part of cell structure
- 5. sugars, starches, cellulose
- broken down as a source of chemical energy; part of cell structure
- 7. fats, oils, phospholipids
- many functions, including movement, transport, chemical catalysts
- 9. enzymes, hemoglobin
- store genetic information, build proteins.

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