

MAKING ESTERS

This page describes ways of making esters in the lab from alcohols and phenols using carboxylic acids, acyl chlorides (acid chlorides) or acid anhydrides as appropriate.

Making esters using carboxylic acids

This method can be used for converting alcohols into esters, but it doesn't work with phenols - compounds where the -OH group is attached directly to a benzene ring. Phenols react with carboxylic acids so slowly that the reaction is unusable for preparation purposes.

The chemistry of the reaction

Esters are produced when carboxylic acids are heated with alcohols in the presence of an acid catalyst. The catalyst is usually concentrated sulphuric acid. Dry hydrogen chloride gas is used in some cases, but these tend to involve aromatic esters (ones where the carboxylic acid contains a benzene ring). If you are a UK A level student you won't have to worry about these.

The esterification reaction is both slow and reversible. The equation for the reaction between an acid RCOOH and an alcohol R'OH (where R and R' can be the same or different) is:



So, for example, if you were making ethyl ethanoate from ethanoic acid and ethanol, the equation would be:



Note: The [mechanism for the esterification reaction](#) is covered in the catalysis section of this site. It is not required for any UK A level (or equivalent) chemistry syllabus.

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Grade12 Experiment Esters

Robert R. Redfield



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Journal of the Society of Chemical Industry Society of Chemical Industry (Great Britain),1920 **Agricultural and Biological Chemistry** ,1981

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