

## QUESTION 5

- 5.1
- 5.1.1 Aufbau diagram ✓✓ (2)
  - 5.1.2 Sulphur ✓✓ (2)
  - 5.1.3  $1s^2 2s^2 2p^4 3s^2 3p^2$  ✓✓ (2)
  - 5.1.4 Group 16 ✓ (2)
  - Period 3 ✓
  - 5.1.5 The element has 6 valence electrons ✓ and belong to the third horizontal row of the periodic table. ✓ (2)
  - 5.1.5 Non-metal ✓ (1)
- 5.2.
- 5.2.1 5 and 1, respectively (1)
  - 5.2.2  $\begin{array}{c} \text{Y} \text{ : } \text{X} \text{ : } \text{Y} \\ \vdots \\ \text{Y} \end{array}$  ✓ (2)
  - 5.2.4 N ✓ and H ✓, or P and H (2)
- [18]

## QUESTION 6

- 6.1
- (a) 24 ✓
  - (b) 12 ✓
  - (c)  $\text{Na}^+$  ✓
  - (d) 10 ✓
  - (e) 12 ✓
  - (f) 17 ✓
  - (g) 19 ✓
- 6.2 Relative atomic mass =  $(0,76 \times 35)$  ✓ +  $(0,24 \times 37)$  ✓  
= 35,48 ✓ (3)
- 6.3 Isotopes ✓✓ (2)
- [12]

## QUESTION 7

- 7.1 Sodium hydroxide ( $\text{NaOH}$ ) ✓  
Hydrogen ✓ or  $\text{H}_2(\text{g})$  (2)
- 7.2  $2\text{Mg}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2\text{MgO}(\text{s})$  ✓✓ (bal) (3)
- 7.3 Sodium + water → sodium hydroxide + water (5)
- $2\text{Na}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{NaOH}(\text{aq}) + \text{H}_2(\text{g})$  ✓ ✓ (Bal)
- 7.4 Because they react vigorously with air to form oxides ✓ and with water to form hydroxides. ✓ (2)
- 7.5 It is an inert colourless gas ✓  
It is unreactive ✓ (2)
- 7.6 Noble gases (1)
- 7.7 Because they don't react, it was not easy to discover them, so no one knew they existed. ✓✓ (2)

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**SB Merriam**



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