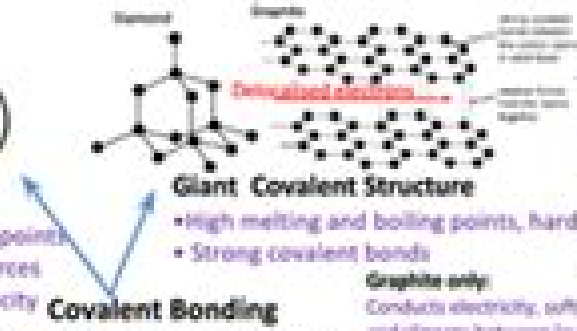
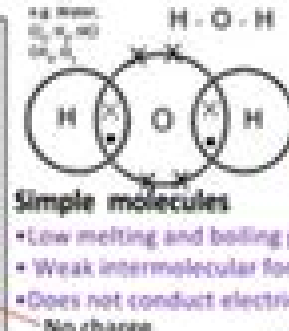




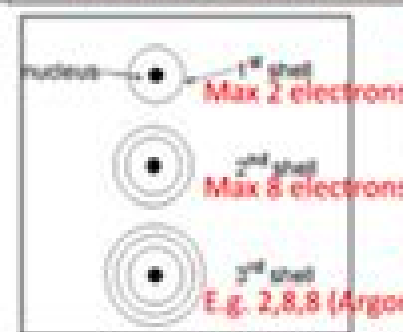
Periodic Table of the Elements

An atom:
Number of Protons = Number of Electrons
Mass number = Protons + Neutrons

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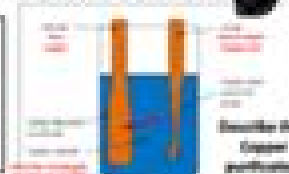
Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
E.g. Lithium 2,1 is a group 1 element.

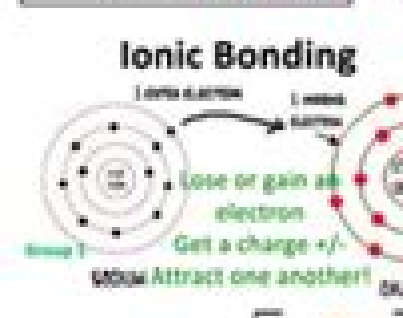
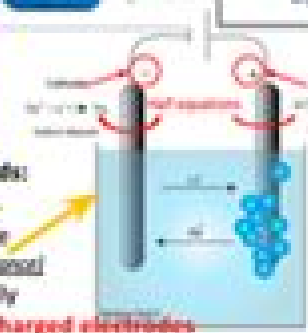


OIL RIG

Oxidation is **lose** (of electrons)
Reduction is **gain** (of electrons)



Electrolysis:
Splitting up a substance using electricity

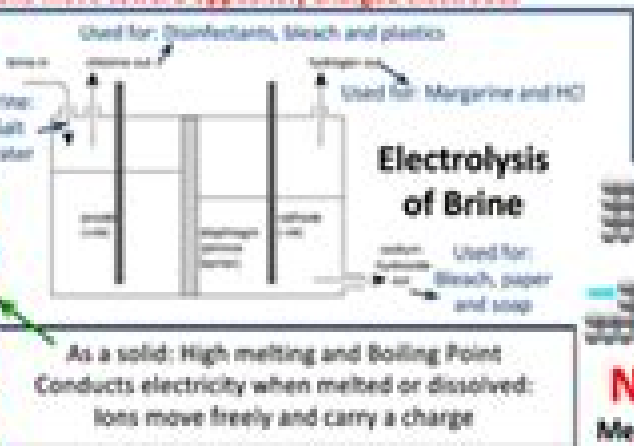
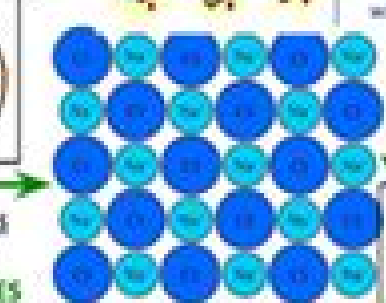


ISOTOPES
are atoms of the same element with different numbers of neutrons

When there are 2 possible products
The less reactive forms at the electrode

Electrolysis needs:

- 2 electrodes
- An electrolyte (ionic or molten substance)
- A power supply



Metal ions only arranged in layers
ELECTROSTATIC ATTRACTION
between
+IONS and ELECTRONS

Metal ions slide over one another making them **MALLEABLE**

Metals conduct **HEAT** well

NANO-TECHNOLOGY
Means **REALLY REALLY REALLY** small!!!!

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