

(b) A student carries out an experiment to measure the enthalpy change of neutralisation.

(i) Define the term *enthalpy change of neutralisation*.

.....

 [1]

(ii) The student measures out 35.0 cm^3 of 2.40 mol dm^{-3} NaOH and 35.0 cm^3 of 2.40 mol dm^{-3} HCl. The temperature of each solution is the same.

The student mixes the two solutions. The temperature rises by 16.5°C .

The specific heat capacity of the mixture is $4.18\text{ J g}^{-1}\text{ K}^{-1}$. Assume that the density of the mixture is 1.00 g cm^{-3} .

Calculate the enthalpy change of neutralisation, in kJ mol^{-1} .

enthalpy change of neutralisation = kJ mol^{-1} [3]

(iii) The student repeats the experiment using 70.0 cm^3 of 1.20 mol dm^{-3} HCl instead of 35.0 cm^3 of 2.40 mol dm^{-3} HCl.

Explain why the temperature rise is less, and predict what the temperature rise will be.

.....

 [2]

[Total: 11]

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