

(ix) The standard potential of the cell in the following reaction is \_\_\_\_\_



(  $E_{\text{Cd}^{2+}}^{\circ} = -0.403 \text{ V}$ ,  $E_{\text{Cu}^{2+}}^{\circ} = 0.334 \text{ V}$  )

(a)  $-0.737 \text{ V}$

(b)  $0.737 \text{ V}$

(c)  $-0.069 \text{ V}$

(d)  $0.069 \text{ V}$

(x) The value of  $[\text{H}_3\text{O}^+]$  in  $\text{mol lit}^{-1}$  of  $0.001 \text{ M}$  acetic acid solution ( $K_a = 1.8 \times 10^{-5}$ ) is \_\_\_\_\_

(a)  $1.34 \times 10^{-1}$

(b)  $1.34 \times 10^{-2}$

(c)  $1.34 \times 10^{-3}$

(d)  $1.34 \times 10^{-4}$

**Q.2. Answer the following questions:**

**[8]**

(i) Write the product formed when alkyl halide reacts with silver nitrite.

(ii) Write the name of product formed, when acetone is treated with 2, 4-dinitrophenyl hydrazine.

(iii) Write the name of biodegradable polyamide copolymer.

(iv) Identify the molecularity of following elementary reaction:



(v) What is the action of selenium on magnesium metal?

(vi) Write the name of isomerism in the following complexes:



(vii) Write the name of the alloy used in Fischer Tropsch process in the synthesis of gasoline.

(viii) Henry's law constant for  $\text{CH}_3\text{Br}_{(\text{g})}$  is  $0.159 \text{ mol dm}^{-3} \text{ bar}^{-1}$  at  $25^\circ\text{C}$ . What is solubility of  $\text{CH}_3\text{Br}_{(\text{g})}$  in water at same temperature and partial pressure of  $0.164 \text{ bar}$ ?

### SECTION – B

**Attempt any EIGHT of the following questions:**

**[16]**

**Q.3.** Explain pseudo-first order reaction with suitable example.

**Q.4.** Write the consequences of Schottky defect with reasons.

**Q.5.** What is the action of following on ethyl bromide:

(i) Na in dry ether

(ii) Mg in dry ether

**Q.6.** Explain formation of peptide linkage in protein with an example.

**Q.7.** Derive an expression to calculate molar mass of non volatile solute by osmotic pressure measurement.

**Q.8.** Explain monodentate and ambidentate ligands with example.

**Q.9.** Explain the trends in the following atomic properties of group 16 elements:

(i) Atomic radii

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**S. K. Patel**



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