

True-False Classify each of the following statements as always true, **AT**; sometimes true, **ST**; or never true, **NT**.

- _____ 1. According to Dalton's atomic theory, atoms are composed of protons, electrons, and neutrons.
- _____ 2. Atoms of elements are electrically neutral.
- _____ 3. The mass of an electron is equal to the mass of a neutron.
- _____ 4. The charge on all protons is the same.
- _____ 5. The atomic number of an element is the sum of the protons and electrons in the atom.
- _____ 6. The atomic number of an element is the whole number that decreases as you read across each row of the periodic table from left to right.
- _____ 7. An atom of nitrogen has 7 protons and 7 neutrons.
- _____ 8. Relative atomic masses are measured in amus.
- _____ 9. The number of neutrons in the nucleus can be calculated by subtracting the atomic from the mass number.

10. Complete the following table.

Element	Symbol	Atomic Number	Mass Number	Number of Protons	Number of Electrons	Number of neutrons
Carbon			12		6	
	K	19				21
		12				12
Helium		2	4			
		5				6

11. Complete the following table

Element	Symbol	Number of Protons	Number of electrons	Number of neutrons	Atomic Number	Mass Number
		25				53
			11	12		
		35		45		
					39	89
			33			75
	Ac					227

12. Fill in the following Table

Element	Symbol	Atomic Number	Mass Number	Number of neutrons
nitrogen-15				8
	$^{22}_{10}\text{Ne}$			
Beryllium-9		4		

13. Use the following information to determine the atomic mass of chlorine. Two isotopes are known: chlorine-35 (mass = 35.0 amu) and chlorine-37 (mass = 37.0 amu). The relative abundance's are 75.4% and 24. 6%, respectively.

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