

Test Unit 3 SCH3U #10

Name: _____

$\frac{\text{K}}{15}$ $\frac{\text{T}}{20}$ $\frac{\text{C}}{5}$ $\frac{\text{A}}{7}$ Assessment of Learning (Product)

Multiple Choice: Circle the most correct answer. (10K)

1. What is the molar mass of AuCl_3 ?

- A) 96 g B) 130 g C) 232.5 g D) 303.6 g

2. Avogadro's number of representative particles is equal to one _____.

- A) kilogram B) gram C) Kelvin D) mole

3. How many moles of tungsten atoms are in 4.8×10^{22} atoms of tungsten?

- A) 8.0×10^2 moles B) 8.0×10^1 moles C) 1.3×10^{-1} moles D) 1.3×10^{-2} moles

4. Determine the mass in grams of 0.0489 mol cobalt.

- A) 0.00812 g B) 0.294 g C) 2.88 g D) 0.000830 g

5. What is the mass in grams of 1.02×10^{24} atoms manganese (Mn)?

- A) 0.112×10^1 g B) 0.169×10^6 g C) 9.30×10^{-1} g D) 9.30×10^6 g

6. Which of the following is/are true?

- i. Samples of water coming from different sources will still have the same percentage composition
- ii. The percentage composition of a compound can ONLY be calculated with knowledge of the compound's formula
- iii. Percent composition is the relative volume of each element in the compound
- iv. Two compounds can have identical percentage compositions

- A) i, ii B) i, ii, iii C) ii, iii D) i, iv E) iv only

7. Which of the following is a candidate for an empirical formula?

- A) $\text{C}_{24}\text{H}_{40}\text{O}_{22}$ B) N_2O_2 C) CH_3OH D) C_6H_6 E) $\text{Na}_2\text{S}_2\text{O}_6$

8. The mass percent of nitrogen, in nitric acid, HNO_3 , is

- A) 1.6% B) 11.1% C) 15.6% D) 22.2%

9. Hydrogen peroxide is known to contain 5.9% hydrogen and 94.1% oxygen, by mass. If its molar mass is 34.0 g/mol, its molecular formula is

- A) HO B) H_2O_2 C) HO_2 D) HNO_2

Grade 11 Stoichiometry Test

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