

I. Multiple Choice Questions (Type-I)

1. On addition of conc. H_2SO_4 to a chloride salt, colourless fumes are evolved but in case of an iodide salt, violet fumes come out. This is because

- (i) H_2SO_4 reduces HI to I_2
- (ii) HI is of violet colour
- (iii) HI gets oxidised to I_2
- (iv) HI changes to HIO_3

Solution:

Option (iii) is the answer.

2. In qualitative analysis when H_2S is passed through an aqueous solution of salt acidified with dil. HCl , a black precipitate is obtained. On boiling the precipitate with dil. HNO_3 , it forms a solution of blue colour. Addition of excess of aqueous solution of ammonia to this solution gives _____.

- (i) a deep blue precipitate of $\text{Cu}(\text{OH})_2$
- (ii) a deep blue solution of $[\text{Cu}(\text{NH}_3)_4]^{2+}$
- (iii) a deep blue solution of $\text{Cu}(\text{NO}_3)_2$
- (iv) a deep blue solution of $\text{Cu}(\text{OH})_2 \cdot \text{Cu}(\text{NO}_3)_2$

Solution:

Option (ii) is the answer.

3. In a cyclotrimetaphosphoric acid molecule, how many single and double bonds are present?

- (i) 3 double bonds; 9 single bonds
- (ii) 6 double bonds; 6 single bonds
- (iii) 3 double bonds; 12 single bonds
- (iv) Zero double bonds; 12 single bonds

Solution:

Option (iii) is the answer.

4. Which of the following elements can be involved in $p\pi-d\pi$ bonding?

- (i) Carbon
- (ii) Nitrogen
- (iii) Phosphorus
- (iv) Boron

Solution:

Option (iii) is the answer.

5. Which of the following pairs of ions are isoelectronic and isostructural?

- (i) CO_3^{2-} , NO_3^-
- (ii) ClO^- , CO_3^{2-}
- (iii) SO_3^{2-} , NO_3^-
- (iv) ClO^- , SO_3^{2-}

Solution:

Option (i) is the answer.

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Christian G. Meyer



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