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Answer Sheet Number. _____

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COMPUTER SCIENCE SSC –I
Section-A (Marks 12)

NOTE: Section –A is compulsory .All parts of this section are to be answered on the question paper itself. It should be completed in the first 20 minutes and handed over to the Centre Superintendent.Deleting /overwriting is not allowed .Do not use lead pencil.

Q.1. Circle the correct option i.e. A/B/C/Each part carries one mark.

I. The data rate of a transmission line that uses modem is 56kbps.What will be the Baud rate.

- A. 72kbps
- B. 126mbps

- C. 56kbps
- D. None of the above

II. Which of the following impairment refer to an undesired signals enter the path of the transmitted signal due to electromagnetic radiations?

- A. Distortion
- B. Cross Talk

- C. Attenuation
- D. Noise

III. Which of the following provide high speed internet connection?

- A. Dial-up connection
- B. ISDN connection

- C. DSL connection
- D. CDMA connection

IV. Which of the following network topology uses a device called terminator?

- A. Ring topology
- B. Mesh topology

- C. Bus topology
- D. Star topology

V. Which of the following authentication method is used for opening security gates?

- A. Username and password
- B. Access card

- C. Personal identification number
- D. Biometrics

VI. What is making legal copies of copyright software for use on other computers or sale called?

Grade 1final Physics Exam Paper 1

Oswaal Editorial Board



Grade 1final Physics Exam Paper 1:

This book delves into Grade 1final Physics Exam Paper 1. Grade 1final Physics Exam Paper 1 is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Grade 1final Physics Exam Paper 1, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Grade 1final Physics Exam Paper 1
 - Chapter 2: Essential Elements of Grade 1final Physics Exam Paper 1
 - Chapter 3: Grade 1final Physics Exam Paper 1 in Everyday Life
 - Chapter 4: Grade 1final Physics Exam Paper 1 in Specific Contexts
 - Chapter 5: Conclusion
2. In chapter 1, this book will provide an overview of Grade 1final Physics Exam Paper 1. This chapter will explore what Grade 1final Physics Exam Paper 1 is, why Grade 1final Physics Exam Paper 1 is vital, and how to effectively learn about Grade 1final Physics Exam Paper 1.
3. In chapter 2, this book will delve into the foundational concepts of Grade 1final Physics Exam Paper 1. The second chapter will elucidate the essential principles that need to be understood to grasp Grade 1final Physics Exam Paper 1 in its entirety.
4. In chapter 3, the author will examine the practical applications of Grade 1final Physics Exam Paper 1 in daily life. The third chapter will showcase real-world examples of how Grade 1final Physics Exam Paper 1 can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Grade 1final Physics Exam Paper 1 in specific contexts. This chapter will explore how Grade 1final Physics Exam Paper 1 is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Grade 1final Physics Exam Paper 1. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Grade 1final Physics Exam Paper 1.

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web k sp of calcium hydroxide 1 author j m mccormick last update ocober 13 2013 introduction an equilibrium constant K is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of K

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product expression describes in mathematical terms the equilibrium

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^{-}$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ 0 M hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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oh 2 m ca in saturated ca oh 2 m ksp of ca oh 2 average ksp of ca oh 2

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web the chemical equation used was $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be $K_{sp} \text{Ca(OH)}_2$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{sp} \text{Ca}_3(\text{PO}_4)_2 = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^2$

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web lab report 8 determinin of calcium hydroxide the purpose of this was to titrat e a saturated ca oh 2 soultion with a standard hcl solution and find the ksp or the oh concentration was f ound and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochlo

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web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for ca was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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