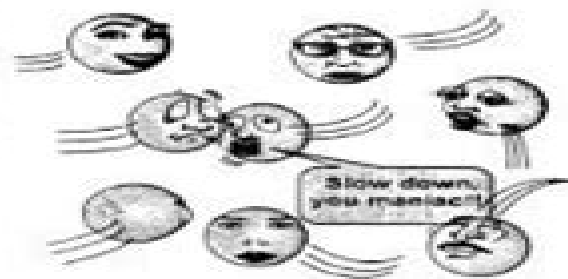


Phase Changes

Heating a substance in a given phase generally causes the temperature to increase. This causes the particles to move faster and collide harder. Harder, faster collisions cause particles to rebound harder moving them further apart. Larger distances between particles weakens the forces of attraction between them. When the forces of attraction are weak enough, the distance between the particles increases markedly and the phase changes. As a result, a solid melts and a liquid evaporates. The reverse happens when a substance cools. As the particles slow down and the collisions weaken, the particles move closer together increasing the forces of attraction between the particles. As a result, a gas condenses and a liquid freezes. At normal atmospheric pressures, some substances change directly from a solid to a gas. This change is called sublimation. This is what occurs when ice cubes shrink in the freezer, or when a wash dries on a clothesline on a cold winter day. Some substances, such as dry ice [$\text{CO}_2(s)$], never pass through a liquid phase at standard pressure [101.3 kPa]. A gas can also change directly to a solid. This is called deposition.



Answer the questions below based on your reading and on your knowledge of chemistry.

1. In the space below are the labels **solid**, **liquid**, and **gas**. Draw arrows between them to represent phase changes. Label each arrow with the name of the phase change it represents. Two examples are done for you. Also, draw an arrow labeled **temperature** to show the direction in which temperature is increasing.



2. Under what conditions does a solid melt? Why? When kinetic energy (temperature) rises above the forces of attraction, a solid melt
3. Under what conditions does a gas condense? Why? When kinetic energy (temperature) decreases and the forces of attraction are felt by the particles

Evan P Silberstein Answer Key

RJ Shavelson



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a break pressure tank which includes a lank structure a pipe arrangement defining a flow line leading through the tank structure from an inlet end connectable to a pipeline from reservoir and entering the tank structure near the top end thereof o an cullet end connectable to a pipeline leading to water supply points and leaving the tank

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construction the break pressure tank consists of the following 1000l 3cr12 stainless steel tank basket strainer with flush valve butterfly valve for isolation diaphragm operated level control valve inlet outlet and overflow pipe work operation the diaphragm valve in the bpt opens and closes as system demand fluctuates

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break pressure tank is a structure that is located between a water reservoir and supply point with the aim of reducing the pressure in the system to zero atmospheric pressure conventional break pressure tank is constructed of concrete in rectangular shape with the depth of the tank about 1.2m

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break pressure tanks which are usually made of concrete or ferrocement if break pressure tanks are used the hydraulic gradient starts again at the tank water level if suitably sized these tanks can be used within the system as storage tanks to meet peak demand reservoir storage tanks although a village's water needs are based upon a

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research and insight break tank break tanks are implemented in boosting systems to supply the system if the mains supply becomes insufficient tanks are also implemented in series connected boosting systems a break tank is also used to ensure that a surge from the starting and stopping pumps doesn't affect the mains distribution

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