

**123****III**

Total No. of Questions – 21

Regd.

Total No. of Printed Pages – 2

No.

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Part – III
CHEMISTRY, Paper – I
(English Version)

Time : 3 Hours]

[Max. Marks : 60

Note : Read the following instructions carefully :

- (1) Answer all the questions of Section – ‘A’. Answer any six questions in Section – ‘B’ and any two questions in Section – ‘C’.
- (2) In Section – ‘A’, questions from Sr. Nos. 1 to 10 are of “Very short answer type”. Each question carries two marks. Every answer may be limited to two or three sentences. Answer all these questions at one place in the same order.
- (3) In Section – ‘B’, questions from Sr. Nos. 11 to 18 are of “Short answer type”. Each question carries four marks. Every answer may be limited to 75 words.
- (4) In Section – ‘C’, questions from Sr. Nos. 19 to 21 are of “Long answer type”. Each question carries eight marks. Every answer may be limited to 300 words.
- (5) Draw labelled diagrams, wherever necessary, for questions in Section – ‘B’ and Section – ‘C’.

SECTION – A**10 × 2 = 20****Note :** Answer all questions :

1. Why KO_2 paramagnetic?
2. Describe the importance of plaster of paris.
3. H_2O has a higher boiling point than that of H_2S . Give reasons.
4. State First Law of thermodynamics.
5. State Graham’s law of diffusion.
6. Calculate the weight of 0.1 mole of Sodium carbonate.

[1 of 2]

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Intermediate 1st Year Chemistry Question Paper 2013

Albert A Gayle



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