

Chemical Reaction Example Guided Practice Problems

Directions: Write balanced chemical equations for each of the following reactions. Be sure to include phases of matter.

1. Solid iron (III) oxide reacts with gaseous hydrogen chloride to form iron (III) chloride and water vapor.
2. Nitric acid dissolved in water reacts with solid sodium carbonate to form liquid water and carbon dioxide gas and sodium nitrate dissolved in water.

Convert these equations to sentences

3. $\text{Fe}_2\text{O}_3 + \text{O}_{2(g)} \rightarrow \text{Fe}_2\text{O}_{3(g)}$
4. $\text{Cu}_{2(s)} + \text{AgNO}_{3(aq)} \rightarrow \text{Ag}_{2(s)} + \text{Cu}(\text{NO}_3)_{2(aq)}$
5. $\text{NO}_{2(g)} \rightarrow \text{N}_{2(g)} + \text{O}_{2(g)}$

Additional Practice

6. Silver nitrate solution and potassium chromate solution, when mixed together, produce solid silver chromate and potassium nitrate solution.
7. Phosphoric acid solution reacts with barium chloride solution to produce solid barium phosphate and hydrochloric acid solution.
8. Solid magnesium burns in air to produce magnesium oxide, a solid white powder.
9. When solid sulfur is exposed to gaseous fluorine, sulfur hexafluoride gas is formed.
10. When exposed to manganese (IV) oxide, hydrogen peroxide rapidly breaks down into water and oxygen gas.
11. Solid sodium carbonate reacts with hydrochloric acid solution to produce carbon dioxide gas, water and sodium chloride solution.

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Table of Contents Guided Practice Problems Chemistry Ch 16

1. Understanding the eBook Guided Practice Problems Chemistry Ch 16
 - The Rise of Digital Reading Guided Practice Problems Chemistry Ch 16
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Practice Problems Chemistry Ch 16
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Practice Problems Chemistry Ch 16
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Practice Problems Chemistry Ch 16
 - Personalized Recommendations
 - Guided Practice Problems Chemistry Ch 16 User Reviews and Ratings
 - Guided Practice Problems Chemistry Ch 16 and Bestseller Lists
5. Accessing Guided Practice Problems Chemistry Ch 16 Free and Paid eBooks
 - Guided Practice Problems Chemistry Ch 16 Public Domain eBooks
 - Guided Practice Problems Chemistry Ch 16 eBook Subscription Services
 - Guided Practice Problems Chemistry Ch 16 Budget-Friendly Options
6. Navigating Guided Practice Problems Chemistry Ch 16 eBook Formats
 - ePub, PDF, MOBI, and More
 - Guided Practice Problems Chemistry Ch 16 Compatibility with Devices
 - Guided Practice Problems Chemistry Ch 16 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guided Practice Problems Chemistry Ch 16
 - Highlighting and Note-Taking Guided Practice Problems Chemistry Ch 16
 - Interactive Elements Guided Practice Problems Chemistry Ch 16

8. Staying Engaged with Guided Practice Problems Chemistry Ch 16
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Practice Problems Chemistry Ch 16
9. Balancing eBooks and Physical Books Guided Practice Problems Chemistry Ch 16
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Practice Problems Chemistry Ch 16
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Guided Practice Problems Chemistry Ch 16
 - Setting Reading Goals Guided Practice Problems Chemistry Ch 16
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Practice Problems Chemistry Ch 16
 - Fact-Checking eBook Content of Guided Practice Problems Chemistry Ch 16
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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