

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

Guided And Study Workbook Answers Covalent Bonding

Jessica J Manson



Guided And Study Workbook Answers Covalent Bonding:

Adopting the Tune of Term: An Psychological Symphony within **Guided And Study Workbook Answers Covalent Bonding**

In a world taken by displays and the ceaseless chatter of fast interaction, the melodic beauty and emotional symphony created by the published term frequently disappear into the background, eclipsed by the constant sound and disruptions that permeate our lives. But, located within the pages of **Guided And Study Workbook Answers Covalent Bonding** a wonderful fictional value filled with natural feelings, lies an immersive symphony waiting to be embraced. Constructed by a masterful musician of language, this interesting masterpiece conducts readers on an emotional journey, well unraveling the hidden tunes and profound influence resonating within each cautiously constructed phrase. Within the depths of this touching evaluation, we can explore the book is central harmonies, analyze their enthralling publishing model, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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