

1/5/16

Mole Ratio Worksheet

Key

① Given this equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, write the following molar ratios:

- a) N_2 / H_2 1:3
 b) N_2 / NH_3 1:2
 c) H_2 / NH_3 3:2

② Given the following equation: $8\text{H}_2 + \text{S}_8 \rightarrow 8\text{H}_2\text{S}$, write the following molar ratios:

- a) $\text{H}_2 / \text{H}_2\text{S}$ 8:8 \rightarrow 1:1 (reduced)
 b) H_2 / S_8 8:1
 c) $\text{H}_2\text{S} / \text{S}_8$ 8:1

③ Answer the following questions for this equation: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

- a) What is the $\text{H}_2 / \text{H}_2\text{O}$ molar ratio? 2:2 \rightarrow 1:1 (reduced)
 b) Suppose you had 20 moles of H_2 on hand and plenty of O_2 , how many moles of H_2O could you make?

$$\frac{20 \text{ mole H}_2}{2 \text{ H}_2} \times \frac{2 \text{ H}_2\text{O}}{2 \text{ H}_2} = \boxed{20 \text{ mole H}_2\text{O}}$$

 c) What is the $\text{O}_2 / \text{H}_2\text{O}$ molar ratio?
 1:2
 d) Suppose you had 20 moles of O_2 and enough H_2 , how many moles of H_2O could you make?

$$\frac{20 \text{ mol O}_2}{1 \text{ O}_2} \times \frac{2 \text{ H}_2\text{O}}{2 \text{ H}_2} = \boxed{40 \text{ mol H}_2\text{O}}$$

4) Use this equation: $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$, for the following problems

- a) If you used 1 mole of N_2 , how many moles of NH_3 could be produced? $\boxed{2}$
 b) If 10 moles of NH_3 were produced, how many moles of N_2 would be required?

$$\frac{10 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{1 \text{ N}_2}{1 \text{ N}_2} = \boxed{5 \text{ mol N}_2}$$

 c) If 3.00 moles of H_2 were used, how many moles of NH_3 would be made?

$$\frac{3 \text{ mol H}_2}{3 \text{ H}_2} \times \frac{2 \text{ NH}_3}{2 \text{ NH}_3} = \boxed{2 \text{ mol NH}_3}$$

 d) If 0.600 moles of NH_3 were produced, how many moles of H_2 are required?

$$\frac{0.6 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{3 \text{ H}_2}{1 \text{ N}_2} = \boxed{0.9 \text{ mol H}_2}$$

Finding Mass Using Mole Ratio Answers

M Lipman

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