

## Reactions in Aqueous Solutions: Metathesis Reactions and Net Ionic Equations

## 9 Pre-lab Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. Write molecular, complete ionic, and net ionic equations for the reactions that occur, if any, when solutions of the following substances are mixed:
  - (a) nitric acid and potassium carbonate
  - (b) sodium bromide and lead nitrate
  - (c) acetic acid and calcium hydroxide
  - (d) calcium nitrate and sodium sulfate
  - (e) ammonium chloride and lithium hydroxide
2. Which of the following are not water-soluble:  $\text{Ba}(\text{NO}_3)_2$ ,  $\text{FeCl}_3$ ,  $\text{CuCO}_3$ ,  $\text{CuSO}_4$ ,  $\text{ZnS}$ , and  $\text{ZnSO}_4$ ?
3. Write equations for the decomposition of  $\text{H}_2\text{CO}_3$  and  $\text{H}_2\text{SO}_3$ .
4. At what temperature (from your graph) do  $\text{KNO}_3$  and  $\text{NaCl}$  have the same molar solubility?

# Metathesis Reaction Lab Answers

**Guillaume Favre**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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