

Guided Reading Chapter 12 Section 1

1. Atoms are made of three basic components. Name these three particles.
2. One of two forces that hold atoms together is called _____.
3. A particle with a negative charge that was discovered by English physicist, _____, is called the _____.
4. Several scientists, including Ernest Rutherford, performed an experiment that helped improve the model of atom by realizing the inner core of the atom housed most of the mass. This core has a special name, the

a) Electron b) nucleus c) proton
5. The two particles that make up most of the mass of the atom are called the _____ and the _____.

a) proton, electron b) neutron, electron c) proton, neutron
6. Which particle occupies the space outside the nucleus, in a "cloud?"
7. The number of _____ is different for each element and is known as the _____ number.
8. The electric charge on an atom is neutral because the number of protons is the same as the number of _____, and each of their individual charges cancel out.
9. _____ are atoms of the same element with a different number of neutrons.

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exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 va
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temperature 2 the beaker now contains sodium chloride and water and the

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web reacting two solutions eg acid and alkali place the polystyrene cup inside the glass beaker to make it more stable
measure an appropriate volume of each liquid eg 25 cm³ 3 place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less
exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of
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web sep 20 2014 i will illustrate this with some examples when an acid is neutralised by an alkali the reaction is exothermic
eg 1 hcl aq + naoh aq → nacl aq + h₂o l for which ΔH = 57 kJ mol⁻¹ eg 2 hNO₃ aq + KOH aq → KNO₃ aq + H₂O l for which ΔH = 57 kJ mol⁻¹

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together to produce one mole of water for example $\text{hcl aq naoh aq nacl aq h}_2\text{o aq}$ you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts

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