

QUESTION 2

- 2.1 The graph of molecular size versus the boiling point is given below.
The letters **A**, **B** and **C** represent the compounds CH_4 , C_2H_6 and C_3H_8 respectively.



- 2.1.1 Define the term *boiling point*. (2)
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)
- 2.2 Consider the two molecules in the table below.

NAME OF SUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT ($^{\circ}\text{C}$)
Ammonia	NH_3	17	- 33
Phosphine	PH_3	x	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine (1)
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

Grade 11 Phycal Science Question Paper

Wenbin Ji



Grade 11 Physical Science Question Paper:

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Table of Contents Grade 11 Phycal Science Question Paper

1. Understanding the eBook Grade 11 Phycal Science Question Paper
 - The Rise of Digital Reading Grade 11 Phycal Science Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Grade 11 Phycal Science Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Grade 11 Phycal Science Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Grade 11 Phycal Science Question Paper
 - Personalized Recommendations
 - Grade 11 Phycal Science Question Paper User Reviews and Ratings
 - Grade 11 Phycal Science Question Paper and Bestseller Lists
5. Accessing Grade 11 Phycal Science Question Paper Free and Paid eBooks
 - Grade 11 Phycal Science Question Paper Public Domain eBooks

- Grade 11 Phycal Science Question Paper eBook Subscription Services
- Grade 11 Phycal Science Question Paper Budget-Friendly Options
- 6. Navigating Grade 11 Phycal Science Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Grade 11 Phycal Science Question Paper Compatibility with Devices
 - Grade 11 Phycal Science Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Grade 11 Phycal Science Question Paper
 - Highlighting and Note-Taking Grade 11 Phycal Science Question Paper
 - Interactive Elements Grade 11 Phycal Science Question Paper
- 8. Staying Engaged with Grade 11 Phycal Science Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Grade 11 Phycal Science Question Paper
- 9. Balancing eBooks and Physical Books Grade 11 Phycal Science Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Grade 11 Phycal Science Question Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Grade 11 Phycal Science Question Paper
 - Setting Reading Goals Grade 11 Phycal Science Question Paper
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade 11 Phycal Science Question Paper
 - Fact-Checking eBook Content of Grade 11 Phycal Science Question Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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