

MILITARY HEALTH SYSTEM MHS GENESIS

Information for Kimbrough Ambulatory Care Center Beneficiaries

COMING TO KIMBROUGH MARCH 25, 2023

MHS GENESIS is the new electronic health record for the Military Health System (MHS). It integrates inpatient and outpatient electronic health records across the continuum of care from the point of injury to the military treatment facility. MHS GENESIS is integral in the provision and coordination of safe, quality care. Kimbrough will transition to MHS GENESIS March 25, 2023.

How does this affect your access to health care?

As we transition to the new electronic health record system, our providers and staff will need time to adjust to the new work flows, procedures, and processes. We slowed down our operations to ensure safe patient care is maintained during this transition. Please expect delays.

During this time, you can expect to experience:

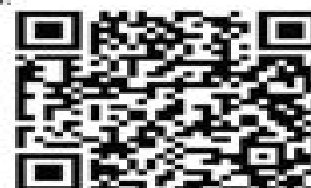
- Longer appointment times
- Increased wait times for Pharmacy pick up
- Reduced appointment availability
- Increased wait times for Lab & Radiology results

How will MHS GENESIS improve your overall experience?

MHS GENESIS allows you to securely access your electronic health records 24/7 and exchange messages with your military health care team when and where it's convenient for your schedule. It provides more efficient management of chronic, complex, and time-sensitive health conditions; provides automated, real-time clinical decision support for health care teams; and increases patient engagement capabilities that allow patients to communicate directly with providers.

How do I access the MHS GENESIS Patient Portal?

To access the MHS GENESIS Patient Portal, visit <https://patientportal.mhsgenesis.health.mil> or scan the QR code.



What can I do within the MHS GENESIS Patient Portal?



- Manage primary care medical and dental appointments
- Review clinical notes, referrals, and lab and test results
- Request prescription renewals
- Exchange secure messages with your health care team
- Monitor health information and view your portal profile
- Complete a pre-visit, dental health history questionnaire online
- Look up high-quality, provider-approved health information related to health issues, lab results, and medications

Patient Support Numbers

- Appt. line: 301-677-8800
- Patient Advocate: 301-677-8836
- Nurse Advice Line: 800-874-2273



For more information, visit health.mil/MHSGENESIS

Genesis Training Manual

S Ashworth



Genesis Training Manual:

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq} + \text{CrO}_4^{2-} \text{aq} \rightarrow \text{Ag}_2\text{CrO}_4 \text{s}$ label 4 2 3

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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^+\text{aq} + 2\text{Cl}^-\text{aq} + \text{Cu}^{2+}\text{aq} + \text{SO}_4^{2-}\text{aq} \rightarrow 2\text{Na}^+\text{aq} + \text{SO}_4^{2-}\text{aq} + \text{CuCl}_2\text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^+\text{aq} + \text{Cl}^-\text{aq} + \text{Na}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{CO}_2\text{g}$ the spectator ions are Cl^- and Na^+
 net ionic equation $\text{H}^+\text{aq} + \text{HCO}_3^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{CO}_2\text{g}$ here is a video on spectator ions
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web the six Na^+ ions and the six Cl^- ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+}\text{aq} + 2\text{PO}_4^{3-}\text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2\text{s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3\text{aq} + \text{KClaq} \rightarrow \text{AgCl s} + \text{KNO}_3\text{aq}$ total ionic $\text{Ag}^+\text{aq} + \text{NO}_3^-\text{aq} + \text{K}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl s} + \text{K}^+\text{aq} + \text{NO}_3^-\text{aq}$ net ionic $\text{Ag}^+\text{aq} + \text{Cl}^-\text{aq} \rightarrow \text{AgCl s}$ 2

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+\text{aq} + \text{OH}^-\text{aq} \rightarrow \text{H}_2\text{O} + \text{l}$ the spectator ions in KOH aq and HNO_3aq are K^+aq and NO_3^-aq how to identify and list the

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loss or gain of one or more electrons ions containing positive charge are cations and those containing negative charge are anions an ion with a net charge of zero is electrically neutral

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