

Electron Configuration Practice

In the space below, write the standard electron configurations of the following elements:

- 1) nitrogen $1s^2 2s^2 2p^3$
- 2) copper $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1 3d^{10}$
- 3) chlorine $1s^2 2s^2 2p^6 3s^2 3p^5$
- 4) calcium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$
- 5) lead $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$

In the space below, write the noble gas notation electron configurations of the following elements:

- 6) barium $[Xe] 6s^2$
- 7) gold $[Xe] 4f^{14} 5d^{10} 6s^1$
- 8) platinum $[Xe] 4f^{14} 5d^9 6s^1$
- 9) xenon $[Kr] 4d^{10} 5s^2 p^6$
- 10) calcium $[Ar] 4s^2$

Determine what elements are denoted by the following electron configurations:

- 11) $1s^2 2s^2 2p^6 3s^2 3p^3$ phosphorous
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$ ruthenium
- 13) $[Kr] 4d^{10} 5s^2 p^5$ iodine
- 14) $[Xe] 4f^{14} 5d^{10} 6s^2$ mercury
- 15) $[Rn] 7s^2$ radium

Determine which of the following electron configurations are not correct:

- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$ $3p^6$ should be a 6
- 17) $1s^2 2s^2 2p^6 3s^2 2p^5$ 1s only holds 2 and there aren't two "2p" orbitals
- 18) $[Cs] 5f^{14} 6d^1 7s^2$ Cs is not a noble gas
- 19) $[Kr] 4d^{10} 5s^2 p^4$ Correct
- 20) $[Ne]$ for Neon Can't use a noble gas to shortcut itself

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