

UNIVERSITY OF PUNE
[4361]-103
F. E. Examination - 2013
ENGINEERING CHEMISTRY
(2012 Pattern)

[Time : 2 Hours]

[Max. Marks : 50]

Instructions :

- (1) Answer Q1 or Q2, Q3 or Q4, Q5 or Q6, Q7 or Q8.
 - (2) Black figures to the right indicate full marks.
 - (3) Neat diagrams must be drawn wherever necessary.
 - (4) Use of logarithmic tables, slide rule, Mollier charts, electronics pocket calculator is allowed.
 - (5) Assume suitable data, if necessary.
-

Q1 a) Explain boiler corrosion and caustic embrittlement as ill effects of using hard water in boilers. State their causes and preventive measures. [6]

b) Which are possible transitions, that occur when molecule absorbs uv- visible radiation? Which type of electronic transitions will be possible in following molecules. [3]



c) State the reference electrode and standard electrode used in pH metery, potentiometry and conductometry. [3]

OR

Q2 a) 50 ml std. hard water containing 1.2 gm CaCO_3 per lit. required 15 ml EDTA solution for the end point. Whereas 50 ml sample water required 19ml of EDTA soln. and 50 ml boiled sample water required 11ml of EDTA soln. for the end point. Calculate, total, temporary and permanent hardness of sample water in ppm. [3]

Engineering Mathematics Pune University Question Papers

Marcel A. Müller



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Unveiling the Magic of Words: A Overview of "**Engineering Mathematics Pune University Question Papers**"

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