

PAPER-I

CHEMISTRY

(SET BY NARAYANA EDUCATIONAL INSTITUTIONS LECTURERS)

- The minimum angular momentum of an electron with the magnetic quantum numbers (-1, 0, +1)
 (1) $\frac{\sqrt{3}h}{\pi}$ (2) $\frac{h}{\pi}$ (3) $\frac{2}{h\pi}$ (4) $\frac{3h}{\pi}$
- Arrange the following molecules in the order of increasing bond angles H_2S , NH_3 , CH_4
 (1) $\text{NH}_3 < \text{H}_2\text{S} < \text{CH}_4$
 (2) $\text{CH}_4 < \text{H}_2\text{S} < \text{NH}_3$
 (3) $\text{H}_2\text{S} < \text{NH}_3 < \text{CH}_4$
 (4) $\text{NH}_3 < \text{CH}_4 < \text{H}_2\text{S}$
- Total kinetic energy of two moles of an Ideal Gas at 127°C is
 (1) 8.31 kJ (2) 9.98 kJ
 (3) 3.20 kJ (4) 5.68 kJ
- Acetic acid undergoes dimerisation in Benzene. Molecular weight of acetic acid in Benzene was observed as 90 in an experiment. Extent of association of acetic acid in the given solvent at the given condition is
 (1) 66% (2) 33% (3) 90% (4) 100%
- Ammonia can be dried over
 (1) Quick lime (2) Slaked lime
 (3) P_2O_5 (4) Conc. H_2SO_4
- $\text{CH}_3\text{CN} + \text{H}_2\text{O} \rightarrow \text{X} + \text{NH}_3$
 The density of 2M solution of X in water is 1.017 gm/ml. The molality of the solution is
 (1) 2 (2) 1.55 (3) 2.26 (4) 1.25
- Electrode potentials of various metals are given
 $\text{Fe} \rightarrow \text{Fe}^{+2} + 2e^-$; $E^\circ = +0.44\text{V}$
 $\text{Pb} \rightarrow \text{Pb}^{+2} + 2e^-$; $E^\circ = +0.13\text{V}$
 $\text{Ag}^+ + e^- \rightarrow \text{Ag}$; $E^\circ = +0.8\text{V}$
 $\text{Cu}^{+2} + 2e^- \rightarrow \text{Cu}$; $E^\circ = +0.34\text{V}$
 Which of the following oxidise Iron?
 (1) Ag only (2) Ag, Cu only
 (3) Ag, Cu, Pb (4) Pb only
- Ethyl alcohol is treated with chlorine. The hybridisation of carbon atoms in final product is
 (1) SP^3, SP^3 (2) SP^3, SP^2
 (3) SP^3, SP (4) SP^2
- α -chlorination of Acetic acid in presence of phosphorous is known as
 (1) Schmidt reaction
 (2) Hell-vohland-Zelensky reaction
 (3) Rosenmund Reaction
 (4) Sandmeyer reaction
- Fe^{+2} is attached to how many nitrogen in haemoglobin
 (1) 4 (2) 5 (3) 6 (4) 3
- The percentage of carbon in CO_2 is
 (1) 27.27% (2) 29.27%
 (3) 30.27% (4) 26.97%
- Formula of Deadburnt Gypsum is
 (1) $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ (2) CaSO_4
 (3) $2\text{CaSO}_4 \cdot \text{H}_2\text{O}$ (4) $2\text{CaSO}_4 \cdot 3\text{H}_2\text{O}$
- In Castner Kellner Process, the Cathode in the middle compartment
 (1) Hg^+ (2) Graphite
 (3) Iron (4) Platinum
- The O-O bond is not present in
 (1) H_2SO_5 (2) CrO_5
 (3) HClO_4 (4) HNO_4
- A mixture of CH_4 and NH_3 heated to 1000°C in the presence of Al_2O_3 can give
 (1) CH_3NH_2 (2) HCN
 (3) $(\text{CH}_3)_2\text{NH}$ (4) CH_3CN
- In carius method halogens are estimated as (X = halogen)
 (1) X_2 (2) BaX_2



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