

## Chapter 15 Alcohols Practice questions

Question number	Answer	Marks	Guidance								
1 (a)	A mixture of liquids is heated to boiling point for a prolonged time  Vapour is formed which escapes from the liquid mixture, is changed back into liquid and returned to the liquid mixture  Any ethanal and ethanol that initially evaporates can then be oxidised	1  1  1									
1 (b)	$\text{CH}_3\text{CH}_2\text{OH} + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{COOH} + 4\text{H}^+ + 4\text{e}^-$	1									
2 (a) (i)	Compounds with the same molecular formula  But different structural formulae / different structures	1  1	This is a definition so should be learnt.								
2 (a) (ii)	$\text{C}_2\text{H}_6\text{O}$ only	1									
2 (a) (iii)	$\text{CH}_2$ only	1									
2 (b)	Potassium dichromate(VI) / $\text{K}_2\text{Cr}_2\text{O}_7$ and acid / acidified / $\text{H}_2\text{SO}_4$ / $\text{H}^+$  Remains orange or no change or no reaction  Orange to green	1  1  1	You can also have $\text{KMnO}_4$ / $\text{H}_2\text{SO}_4$ , but not $\text{HCl}$ .  Or remain purple if $\text{KMnO}_4$ used.  Or goes from purple to colourless if $\text{KMnO}_4$ in acid used or gives brown precipitate or goes green if $\text{KMnO}_4$ neutral or in alkali.								
2 (c)	<div><div><u>Choice of reagents</u></div><table><tr><td>potassium dichromate (VI) / <math>\text{K}_2\text{Cr}_2\text{O}_7</math> and acid / acidified / <math>\text{H}_2\text{SO}_4</math> / <math>\text{H}^+</math> or <math>\text{KMnO}_4</math> / <math>\text{H}_2\text{SO}_4</math></td><td>Fehling's / Benedict's reagent</td><td>Tollens' reagent or <math>\text{AgNO}_3</math> / <math>\text{NH}_3</math> or ammoniacal silver nitrate Not <math>\text{AgNO}_3</math> alone.</td></tr></table><div><u>with the aldehyde C</u></div><table><tr><td>goes orange to green  goes purple to colourless / brown ppt. / green solution if <math>\text{KMnO}_4</math> used red solid  Potassium dichromate is more usual to do in school.</td><td>red solid</td><td>silver mirror</td></tr></table><div><u>with the ketone D</u></div><table><tr><td>remains orange or no change or no reaction or purple for <math>\text{KMnO}_4</math></td><td>remains blue or no change or no reaction</td><td>remains colourless or no change or no reaction</td></tr></table></div> <div>1  1  1</div> <td>These tests always come up, so learn the tests to distinguish aldehydes and ketones. If you can't learn them all pick one to really learn.</td>	potassium dichromate (VI) / $\text{K}_2\text{Cr}_2\text{O}_7$ and acid / acidified / $\text{H}_2\text{SO}_4$ / $\text{H}^+$ or $\text{KMnO}_4$ / $\text{H}_2\text{SO}_4$	Fehling's / Benedict's reagent	Tollens' reagent or $\text{AgNO}_3$ / $\text{NH}_3$ or ammoniacal silver nitrate Not $\text{AgNO}_3$ alone.	goes orange to green  goes purple to colourless / brown ppt. / green solution if $\text{KMnO}_4$ used red solid  Potassium dichromate is more usual to do in school.	red solid	silver mirror	remains orange or no change or no reaction or purple for $\text{KMnO}_4$	remains blue or no change or no reaction	remains colourless or no change or no reaction	These tests always come up, so learn the tests to distinguish aldehydes and ketones. If you can't learn them all pick one to really learn.
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