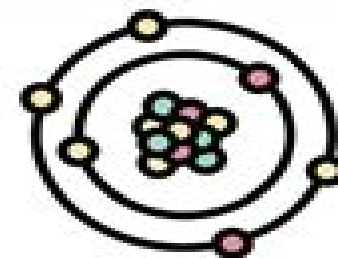


ATOMS



GUIDED READING & PRACTICE WORKSHEETS

NAME _____ **PER** _____ **DATE** _____

ATOMS

1 ELECTRON SHELLS

(1) An electron shell also known as an energy level, is a region around the nucleus in which electrons are likely to be found. Electrons in an atom are arranged in shells around the nucleus. The number of shells an atom has is equal to the number of protons in the atom.

(2) The electrons in an atom fill the lowest energy level first, before moving to the next energy level. This is why the outermost shell of an atom is the least stable and the innermost shell is the most stable. The number of electrons in each shell is limited by the number of protons in the atom.

(3) Some elements can have different numbers of electron shells. This is because the number of shells an atom has is determined by the number of protons in the atom. For example, hydrogen has one shell, while helium has two shells.

(4) The number of electrons in an atom is equal to the number of protons in the atom. This is because the atom is neutral and the number of electrons must equal the number of protons to maintain a neutral charge.

ATOMS PRACTICE

Complete the following table for each element.

Element	Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons
Hydrogen	H	1	1	1	0	1
Helium	He	2	4	2	2	2
Lithium	Li	3	7	3	4	3
Boron	B	5	11	5	6	5
Carbon	C	6	12	6	6	6
Nitrogen	N	7	14	7	7	7
Oxygen	O	8	16	8	8	8
Fluorine	F	9	19	9	10	9
Neon	Ne	10	20	10	10	10

PERIODIC TABLE OF ELEMENTS

Hydrogen (H) is the first element in the periodic table. It has an atomic number of 1 and a mass number of 1. It is a gas at room temperature and is used in many industrial processes.

The periodic table is organized into groups and periods. Groups are vertical columns of elements, and periods are horizontal rows of elements. Elements in the same group have similar chemical properties, while elements in the same period have the same number of electron shells.

The periodic table is a fundamental tool for understanding the properties and behavior of elements. It is used by scientists and students alike to study the periodic law and the periodic table of elements.

GRADES
8-12



NAME _____ **PER** _____ **DATE** _____

ATOMS

1 PARTS OF AN ATOM

(1) Atoms are the basic units of matter and the smallest particles of an element that retain the chemical properties of that element. A particular element is composed of only one type of atom.

(2) Atoms are made up of three types of subatomic particles: protons, neutrons, and electrons. Protons are positively charged particles that are located in the nucleus of an atom. Neutrons are neutral particles and are also located in the nucleus. Electrons are negatively charged particles that orbit the nucleus in electron shells or energy levels. Typically, the number of protons in an atom is equal to the number of electrons.

Label the atom diagram and fill in the table below.

Diagram of an atom showing the nucleus (protons and neutrons) and electron shells (electrons).

Subatomic Particle	Symbol	Charge	Location
Proton	p^+	Positive	Nucleus
Neutron	n^0	Neutral	Nucleus
Electron	e^-	Negative	Electron Shell

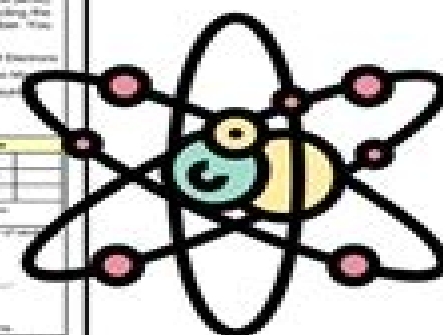
2 ATOMIC NUMBER & MASS

(1) The atomic number is the number of protons in the nucleus of an atom. The number of protons in an atom determines what the element is. For example, only atoms with one proton is hydrogen, while any atom with six protons is carbon.

(2) The mass number is the sum of the number of protons and neutrons in an atom's nucleus. It is measured in atomic mass units (amu). You can calculate the number of neutrons in an atom by subtracting the number of protons in an atom from the atom's mass number. You may round the number to the nearest whole number.

Complete the following table for each element.

Element	Symbol	Atomic Number	Mass Number	Protons	Neutrons	Electrons
Hydrogen	H	1	1	1	0	1
Helium	He	2	4	2	2	2
Lithium	Li	3	7	3	4	3
Boron	B	5	11	5	6	5
Carbon	C	6	12	6	6	6
Nitrogen	N	7	14	7	7	7
Oxygen	O	8	16	8	8	8
Fluorine	F	9	19	9	10	9
Neon	Ne	10	20	10	10	10



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language and vocabulary Students will label each part of the atom learn what compounds are and explore the patterns in the periodic table of elements to find calcium Ca chlorine Cl and helium He through hands on activities These and more science concepts are presented in a way that makes them more accessible to students and easier to understand Written to grade and using simplified language and vocabulary and comprised of reading passages student activities crossword word search comprehension quiz and color mini posters our resource can be used effectively for test prep and your whole class All of our content is aligned to your State Standards and are written to Bloom s Taxonomy and STEM initiatives

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Survival Guide to General Chemistry Patrick E. McMahon,Rosemary McMahon,Bohdan Khomtchouk,2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available

texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction The Official ACT Prep Guide 2019-2020, (Book + 5 Practice Tests + Bonus Online Content) ACT,2019-04-16 The only guide from the makers of the ACT exam packed with 5 genuine full length practice tests and 400 additional questions online This new edition includes A NEW never before seen full length practice test with optional writing test 215 questions 400 online questions that can be filtered and organized into custom practice sets Updated writing prompts and directions Real ACT test forms used in previous years The Official ACT Prep Guide 2019 2020 is the only guide from the makers of the exam and includes actual ACT test forms taken from past ACT exams This updated edition includes 5 actual ACT tests all with optional writing test to help you practice at your own pace and discover areas where you may need more work The Official ACT Prep Guide 2019 2020 provides detailed explanations for every answer and practical tips on how to boost your score on the English math reading science and optional writing tests You ll also get access to special online bonus content developed with the test taking experience in mind Practice with 400 additional test questions that can be organized filtered and tracked for performance Take a closer look at test day learn what to expect and get familiar with the test taking strategies that are right for you The Official ACT Prep Guide 2019 2020 is your definitive guide to getting ready for the ACT and feeling confident and comfortable on test day Differentiating Science Instruction and Assessment for Learners With Special Needs, K□8 Kevin D. Finson,Christine K. Ormsbee,Mary M. Jensen,2011-03-02 Field tested strategies for teaching science to students with special needs Teachers are required to provide appropriate science instruction to all students including children with special needs However they are often left on their own to figure out how to

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Structure of Multielectron Atoms Anders Kastberg, 2020-04-23 This book is about the structure of multielectron atoms and predominantly adopts a perturbative approach to the total Hamiltonian A key concept is the central field approximation and beyond the standard LS coupling and jj coupling schemes intermediate cases are also treated After that the book covers hyperfine structure and other nuclear effects as well as interactions with static external fields Throughout the book an analytical approach is adopted Working knowledge of basic quantum mechanics including the non relativistic hydrogen atom basic angular momentum and perturbation theory is assumed and it begins with a brief recap of the hydrogen orbitals before turning towards the symmetry aspects of multi electron atoms spin orbit interaction and couplings of angular momenta

A Student's Guide Through the Great Physics Texts Kerry Kuehn, 2015-12-17 This book provides a chronological introduction to modern atomic theory which represented an attempt to reconcile the ancient doctrine of atomism with careful experiments performed during the 19th century on the flow of heat through substances and across empty space Included herein are selections from classic texts such as Carnot's Reflection on the Motive Power of Fire Clausius Mechanical Theory of Heat Rutherford's Nuclear Constitution of Atoms Planck's Atomic Theory of Matter and Heisenberg's Copenhagen Interpretation of Quantum Theory Each chapter begins with a short introduction followed by a reading selection Carefully crafted study questions draw out key points in the text and focus the reader's attention on the author's methods analysis and conclusions Numerical and laboratory exercises at the end of each chapter test the reader's ability to understand and apply key concepts from the text Heat Radiation and Quanta is the last of four volumes in A Student's Guide through the Great Physics Texts The book comes from a four semester undergraduate physics curriculum designed to encourage a critical and circumspect approach to natural science while at the same time preparing students for advanced coursework in physics This book is particularly suitable as a college level textbook for students of the natural sciences history or philosophy It might also serve as a textbook for advanced high school or home schooled students or as a thematically organized source book for scholars and motivated lay readers In studying the classic

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