

Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $[Rn] 7s^2 6d^1 5f^{14}$
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ osmium
- 15) $[Rn] 7s^2 5f^{11}$ einsteinium note: no d's, so the p. table referenced has Ac as 5f
- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (take a look at "4d"; must be 3 d not 4!)
- 17) $1s^2 2s^2 2p^6 3s^3 3d^5$ not valid (3p comes after 3s); also no 3 electrons in s
- 18) $[Ra] 7s^2 5f^8$ not valid (radium isn't a noble gas)
- 19) $[Kr] 5s^2 4d^{10} 5p^5$ valid
- 20) $[Xe]$ not valid (an element can't be its own electron configuration)

Electron Configuration Worksheet Answer Key Section 5

3

WJ Hussar



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Table of Contents Electron Configuration Worksheet Answer Key Section 5 3

1. Understanding the eBook Electron Configuration Worksheet Answer Key Section 5 3
 - The Rise of Digital Reading Electron Configuration Worksheet Answer Key Section 5 3
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Worksheet Answer Key Section 5 3
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Worksheet Answer Key Section 5 3
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Worksheet Answer Key Section 5 3

- Personalized Recommendations
- Electron Configuration Worksheet Answer Key Section 5 3 User Reviews and Ratings
- Electron Configuration Worksheet Answer Key Section 5 3 and Bestseller Lists
- 5. Accessing Electron Configuration Worksheet Answer Key Section 5 3 Free and Paid eBooks
 - Electron Configuration Worksheet Answer Key Section 5 3 Public Domain eBooks
 - Electron Configuration Worksheet Answer Key Section 5 3 eBook Subscription Services
 - Electron Configuration Worksheet Answer Key Section 5 3 Budget-Friendly Options
- 6. Navigating Electron Configuration Worksheet Answer Key Section 5 3 eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configuration Worksheet Answer Key Section 5 3 Compatibility with Devices
 - Electron Configuration Worksheet Answer Key Section 5 3 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configuration Worksheet Answer Key Section 5 3
 - Highlighting and Note-Taking Electron Configuration Worksheet Answer Key Section 5 3
 - Interactive Elements Electron Configuration Worksheet Answer Key Section 5 3
- 8. Staying Engaged with Electron Configuration Worksheet Answer Key Section 5 3
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Worksheet Answer Key Section 5 3
- 9. Balancing eBooks and Physical Books Electron Configuration Worksheet Answer Key Section 5 3
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Worksheet Answer Key Section 5 3
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Worksheet Answer Key Section 5 3
 - Setting Reading Goals Electron Configuration Worksheet Answer Key Section 5 3
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Worksheet Answer Key Section 5 3

- Fact-Checking eBook Content of Electron Configuration Worksheet Answer Key Section 5 3
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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