

Electron Configuration Practice
Chemistry

Name: _____

Due Date: _____

How to write an electron configuration:

- Determine the total number of electrons to be represented.
- Use the Aufbau principle to fill the orbitals with electrons for elements 1-23. Refer to electron configuration periodic table for elements after 23
- The sum of the superscripts should equal the total number of electrons. Example:
 $12\text{Mg } 1s^2 2s^2 2p^6 3s^2$

I. Configuration Writing Practice

- Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled.

1. Na: $1s^2 2s^2 2p^6 3s^1$

2. Pb: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^2$

3. Sr: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$

4. U: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^3 6d^1$

5. N: $1s^2 2s^2 2p^3$

* 6. Ag: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^{10}$

7. Ti: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^2$

8. Ce: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^1 5d^1$

9. Cl: $1s^2 2s^2 2p^6 3s^2 3p^5$

10. Hg: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$

Electron Configurations Packet With Answers

Marco Cascella



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Table of Contents Electron Configurations Packet With Answers

1. Understanding the eBook Electron Configurations Packet With Answers
 - The Rise of Digital Reading Electron Configurations Packet With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configurations Packet With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configurations Packet With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configurations Packet With Answers
 - Personalized Recommendations
 - Electron Configurations Packet With Answers User Reviews and Ratings
 - Electron Configurations Packet With Answers and Bestseller Lists
5. Accessing Electron Configurations Packet With Answers Free and Paid eBooks
 - Electron Configurations Packet With Answers Public Domain eBooks
 - Electron Configurations Packet With Answers eBook Subscription Services
 - Electron Configurations Packet With Answers Budget-Friendly Options

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 - Electron Configurations Packet With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configurations Packet With Answers
 - Highlighting and Note-Taking Electron Configurations Packet With Answers
 - Interactive Elements Electron Configurations Packet With Answers
8. Staying Engaged with Electron Configurations Packet With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configurations Packet With Answers
9. Balancing eBooks and Physical Books Electron Configurations Packet With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configurations Packet With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electron Configurations Packet With Answers
 - Setting Reading Goals Electron Configurations Packet With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configurations Packet With Answers
 - Fact-Checking eBook Content of Electron Configurations Packet With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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