

Mean Absolute Deviation

Find the mean absolute deviation of each data set. Round your answers to two decimal places.

1) 12, 19, 10, 18, 21

Data	Mean	Difference	Absolute Value
Sum			

Mean Absolute Deviation = _____

3) 50, 48, 36, 45, 33, 37

Data	Mean	Difference	Absolute Value
Sum			

Mean Absolute Deviation = _____

2) 7, 14, 11, 13, 4, 20

Data	Mean	Difference	Absolute Value
Sum			

Mean Absolute Deviation = _____

4) 8, 5, 2, 9

Data	Mean	Difference	Absolute Value
Sum			

Mean Absolute Deviation = _____

Mean Average Deviation 7th Grade

S Ben Porath

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + \text{O}_2(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ}$ mole
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ}$ mole
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ}$ mole

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 hess s law example 12.1 determine Δh for the target reaction
 $\text{C}_2\text{H}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4\text{O}(\text{g})$
 given the following information

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 $\text{Mg}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{MgO}(\text{s})$ we used

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 cohen 2016 the equations for the reactions in the experiment done are as follows
 $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) + \text{OH}(\text{aq})$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta h = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta h_{\text{products}} > \Delta h_{\text{reactants}}$

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