



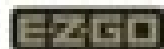
Introducing the  
**Powerwise™ QE** 36 Volt Battery Charger

## MAINTENANCE INSTRUCTIONS

1. For flooded lead-acid batteries, regularly check water levels of each battery cell after charging and add distilled water as required to level specified for battery manufacturer. Follow the safety instructions recommended by the battery manufacturer.
2. Make sure charger connections to battery terminals are tight and clean. Check for deformation/cracks in plastic parts. Check the charger harness for chaffing and rubbing. Inspect all wiring for fraying, loose terminals, chaffing, corrosion or deterioration of insulation.
3. Keep cooling fins free of contamination and do not expose charger to oil, dirt, mud or to direct heavy water spraying when cleaning equipment.
4. Inspect the plug of the battery charger and vehicle receptacle housing for dirt or debris. Monthly, or sooner if needed, clean the DC connector auxiliary contact. To clean the auxiliary contact, slide an entry board between main contact and auxiliary contact located in the hole of the charger plug nearest the rounded corners. (See Cleaning Contact Figures). Press entry board down to apply pressure to the auxiliary contact and slide board in and out of plug approximately 10 to 20 times, keeping pressure applied to the auxiliary contact surface. Repeat cleaning process if charger turns off while the handle is moved back and forth. If, after re-cleaning and retesting, the charger still turns off, check plug for a broken set wire or (X) cord.



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web jan 30 2023 hess s law is due to enthalpy being a state function which allows us to calculate the overall change in enthalpy by simply summing up the changes for each step of the way until product is formed all steps have to proceed at the same temperature and the equations for the individual steps must balance out

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subsequent  $\Delta h$  values  $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g})$

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web chemistry 120 hess s law worksheet calculate  $\Delta h$  for the reaction  $\text{C}_2\text{H}_4(\text{g}) + 2 \text{C}_2\text{H}_6(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) + 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ}$  mole  $\text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ}$  mole  $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ}$

mole

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{NaOH}$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}}$

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