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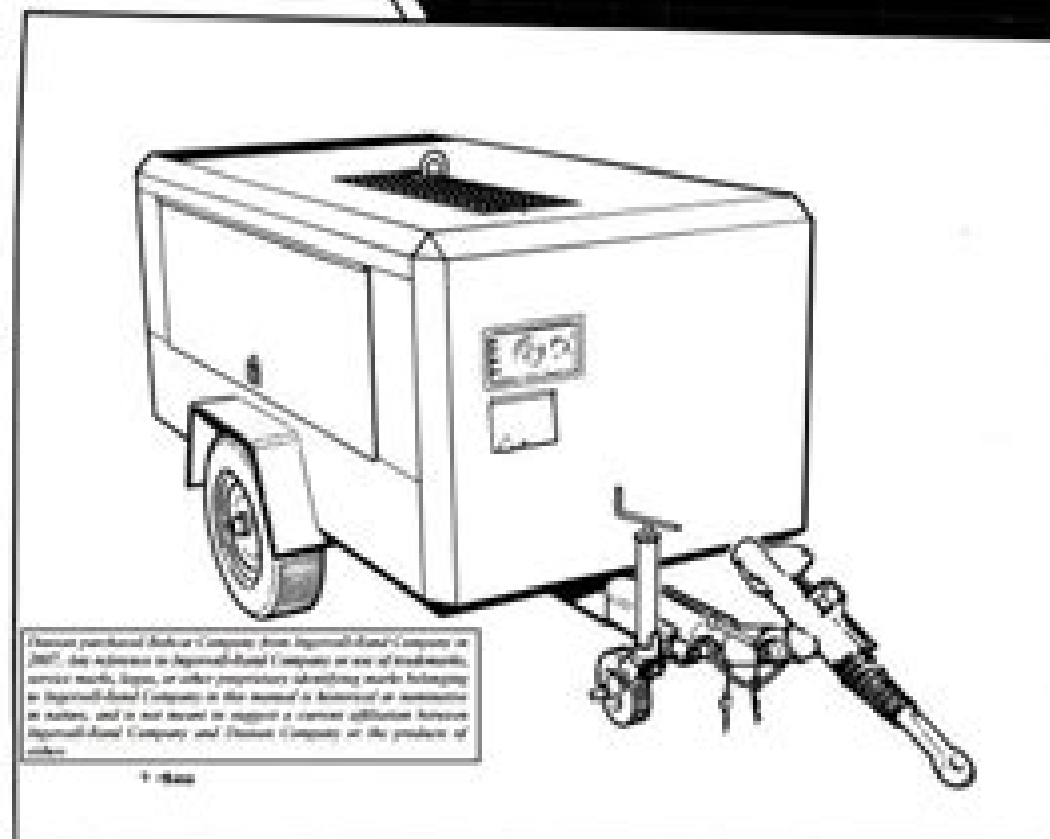
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 $\text{C}_2\text{H}_2 \text{ g} + 2 \text{ H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g} \quad \Delta h = -94.5 \text{ kJ}$
 $\text{H}_2\text{O g} + \frac{1}{2} \text{O}_2 \text{ g} \rightarrow \text{H}_2\text{O l} \quad \Delta h = -71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O g} \quad \Delta h = -283 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \quad \Delta h = 130 \text{ kJ}$

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 $\text{C}_2\text{H}_6 \text{ g} + \frac{7}{2} \text{O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2\text{O l} \quad \Delta h = -1560 \text{ kJ}$
 $\text{H}_2 \text{ g} + \frac{1}{2} \text{O}_2 \text{ g} \rightarrow \text{H}_2\text{O l} \quad \Delta h = -285 \text{ kJ}$

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