

1. True or False?

- C2 is more electron-deficient than C1:



**False.** There are three partially positive carbons in this conjugated molecule. Be sure to draw out the most reasonable resonance structures that place a (+) charge on C (below).



- Of the two terminal diene carbons, **B** is more electron-rich than **A**:



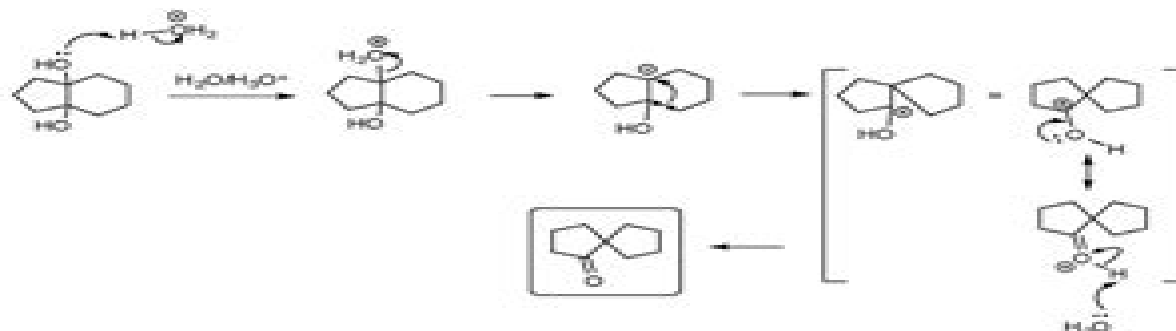
**False.** The O lone pair can delocalize its lone pair into the diene structure. It is possible to place a fully negative charge on the C labeled A.



Below I have drawn the arrow-formalism converting the given Lewis structure to that with a full negative charge on A.



2. Provide the complete arrow-formalism mechanism for the following transformation.



# Gpb Chem 203 Test Unit 2 Answers

**Joacim Rocklöv**



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