

100) Carbon tetrachloride consists of one carbon and four chlorine atoms. How many chlorine atoms are present in 1.00 mole of CCl_4 molecules?

$$\text{one mole of } \text{CCl}_4 \text{ contains } 4 \text{ moles of } \text{Cl} \text{ atoms}$$

$$1 \text{ mole } \text{CCl}_4 \times \frac{4 \text{ moles Cl}}{1 \text{ mole } \text{CCl}_4} = 4 \text{ moles Cl}$$

$$\text{answer} = 4 \text{ moles Cl} \quad \text{100}$$

101) There are 3 monatomic gases in each buffer gas used for plasma.

101) Describe how each of these gases is used.

Helium is used for cooling the plasma.
 Argon is used for shielding the plasma.
 Neon is used for the plasma.

102) Describe how each of these gases is used in the plasma.

Helium is used for cooling the plasma.
 Argon is used for shielding the plasma.
 Neon is used for the plasma.

103) Carbon tetrachloride is a compound of carbon and chlorine. How many chlorine atoms are present in 1.00 mole of CCl_4 molecules?

103) Describe how each of these gases is used in the plasma.

104) What is the empirical formula of the compound formed by the reaction of 1.00 mole of CCl_4 with 1.00 mole of H_2 ?



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