

15. (a) The azeotrope of the benzene – cyclohexane system has a composition of 53.2 mol per cent benzene with a boiling point of 350.6 K at 101.3 kPa. At this temperature, the vapour pressure of benzene (i) is 100.59 kPa and that of cyclohexane (ii) is 99.27 kPa. Determine the activity coefficients for the solution containing 10 mol per cent benzene.

Or

- (b) Estimate the standard free energy change and equilibrium constant at 700 K for the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ given that the standard heat of formation and standard free energy of formation of ammonia at 298 K to be $-46, 100 \text{ J/mol}$ and $-16,500 \text{ J/mol}$ respectively. The specific heat (J/mol K) data are given as a function of temperature (K)

$$C_p = 27.27 + 4.93 \times 10^{-3}T \text{ for } \text{N}_2$$

$$C_p = 27.01 + 3.51 \times 10^{-3}T \text{ for } \text{H}_2$$

$$C_p = 29.75 + 25.11 \times 10^{-3}T \text{ for } \text{NH}_3$$

PART C — (1 × 15 = 15 marks)

16. (a) Differentiate between state function and path function. Prove that internal energy is a state function and work is a path function.

Or

- (b) Applying the method of partial derivatives, prove that the joule Thomson coefficient of an ideal gas is zero.

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