

Chapter 5 Chemistry: Electrons in Atoms**Matching***Match each item with the correct statement below.*

- | | |
|------------------------------|---------------------------|
| a. Pauli exclusion principle | e. electron configuration |
| b. aufbau principle | d. ground state |

- _____ 1. each orbital has at most two electrons
- _____ 2. lowest energy level
- _____ 3. tendency of electrons to enter orbitals of lowest energy first
- _____ 4. arrangement of electrons around atomic nucleus

Multiple Choice*Identify the choice that best completes the statement or answers the question.*

- _____ 5. What is the maximum number of electrons that can occupy the $n = 3$ shell?
- | | | | | |
|-------|------|------|-------|-------|
| a. 50 | b. 4 | c. 8 | d. 32 | e. 18 |
|-------|------|------|-------|-------|
- _____ 6. Which of the following elements is a d -block element?
- | | | | | |
|-------|-------|------|-------|-------|
| a. Cd | b. Cf | c. C | d. Cs | e. Cl |
|-------|-------|------|-------|-------|
- _____ 7. Which element has the electron configuration $1s^2 2s^2 2p^6 3d^6 4s^2$?
- | | | | | |
|-------|-------|-------|-------|-------|
| a. Ag | b. Cd | c. Zn | d. Cu | e. Ge |
|-------|-------|-------|-------|-------|
- _____ 8. Which element has the following ground state electron configuration?
- [Ar] $\boxed{\uparrow\downarrow}\boxed{\uparrow\downarrow}\boxed{\uparrow}\boxed{\uparrow}\boxed{\uparrow}\boxed{\uparrow\downarrow}$
- $3d$ $4s$
- | | | | | |
|-------|-------|-------|-------|-------|
| a. Fe | b. Co | c. Cu | d. Ni | e. Zn |
|-------|-------|-------|-------|-------|
- _____ 9. The $n =$ _____ shell is the lowest that may contain d -orbitals.
- | | | | | |
|------|------|------|------|------|
| a. 2 | b. 4 | c. 3 | d. 5 | e. 6 |
|------|------|------|------|------|
- _____ 10. In Bohr's model of the atom, where are the electrons and protons located?
- | |
|--|
| a. The electrons move around the protons, which are at the center of the atom. |
| b. The electrons and protons move throughout the atom. |
| c. The electrons occupy fixed positions around the protons, which are at the center of the atom. |
| d. The electrons and protons are located throughout the atom, but they are not free to move. |
- _____ 11. How many energy sublevels are in the second principal energy level?
- | | |
|------|------|
| a. 1 | c. 3 |
| b. 2 | d. 4 |
- _____ 12. What is the maximum number of orbitals in the p sublevel?
- | | |
|------|------|
| a. 2 | c. 4 |
| b. 3 | d. 5 |

Electrons In Atoms Chapter Test A

G Psacharopoulos



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