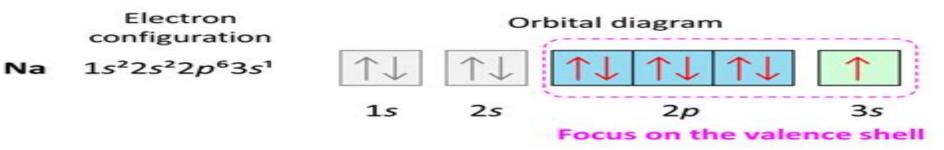
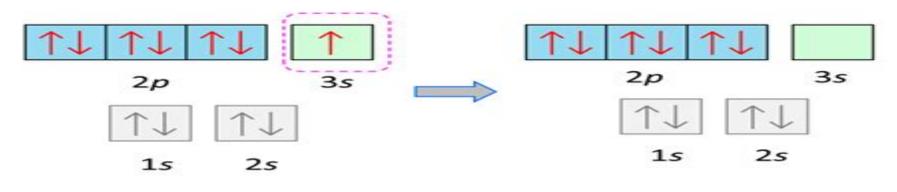
Electron Configuration of Main Group Cations



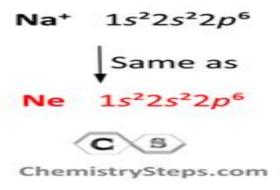
The charge is +1, therefore, one electron is lost:

The electron must be from the outermost valence orbital $-3s^1$



Na $1s^2 2s^2 2p^6 3s^1$

Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are *isoelectronic* with the previous noble gas.



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according to the themes novel pedagogies dynamic teaching environments new approaches in assessment and professional skills each of which is of substantial current interest to the science education communities Providing an overview of contemporary practice this book helps improve student learning outcomes Many of the teaching strategies presented are transferable to other disciplines and are of great interest to the global community of tertiary chemistry educators as well as readers in the areas of secondary STEM education and other disciplines **Chemistry Class - XII Model Paper Chapter** wise Question Answer With Marking Scheme 2023- SBPD Publications SBPD Editorial Board, 2022-10-14 Content 1 Solid State 2 Solution 3 Electrochemistry 4 Chemical Kinetics 5 Surface Chemistry 6 General Principles and Processes of Isolation of Element 7 P Block elements 8 d and f Block Elements 9 Coordination Compunds 10 Haloalkanes and Haloarences 11 Alcohols Phenols and Ethers 12 Aldehydes Ketones and Carboxylic Acid 13 Organic Compounds Containing Nitrogen 14 Biomolecules 15 polymers 16 Chemistry in Everyday life Model Paper Set 1 4 BSEB With OMR Sheet Board Examination AOA AS Chemistry: Unit 1 Foundation Chemistry Student Unit Guide Ebook Papers BSEB CBSE With OMR Sheet Margaret Cross, 2008-09-26 Student Unit Guides are perfect for revision Each guide is written by an examiner and explains the unit requirements summarises the relevant unit content and includes a series of specimen questions and answers There are three sections to each guide Introduction includes advice on how to use the guide an explanation of the skills being tested by the assessment objectives an outline of the unit or module and depending on the unit suggestions for how to revise effectively and prepare for the examination questions Content Guidance provides an examiner's overview of the module s key terms and concepts and identifies opportunities to exhibit the skills required by the unit It is designed to help students to structure their revision and make them aware of the concepts they need to understand the exam and how they might analyse and evaluate topics Question and Answers sample questions and with graded answers which have been carefully written to reflect the style of the unit All responses are accompanied by commentaries which highlight their respective strengths and **Chemical Principles** Peter Atkins, Loretta weaknesses giving students an insight into the mind of the examiner Jones, 2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry s frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text Instructors have the option of whether to incorporate calculus in the coverage of topics The multimedia integration of Chemical Principles is more deeply established than any other text for this course Through the unique eBook the comprehensive Chemistry Portal Living Graph icons that connect the text to the Web and a complete set of animations students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding Shriver and Atkins' Inorganic Chemistry

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in the indian state of uttarakhand

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