

Section 8.5

Objectives

- **Describe** how electronegativity is used to determine bond type.
- **Compare and contrast** polar and nonpolar covalent bonds and polar and nonpolar molecules.
- **Generalize** about the characteristics of covalently bonded compounds.

Review Vocabulary

electronegativity: the relative ability of an atom to attract electrons in a chemical bond

New Vocabulary

polar covalent bond

Electronegativity and Polarity

MAIN Idea A chemical bond's character is related to each atom's attraction for the electrons in the bond.

Real-World Reading Link The stronger you are, the more easily you can do pull-ups. Just as people have different abilities for doing pull-ups, atoms in chemical bonds have different abilities to attract (pull) electrons.

Electron Affinity, Electronegativity, and Bond Character

The type of bond formed during a reaction is related to each atom's attraction for electrons. Electron affinity is a measure of the tendency of an atom to accept an electron. Excluding noble gases, electron affinity increases with increasing atomic number within a period and decreases with increasing atomic number within a group. The scale of electronegativities—shown in **Figure 8.20**—allows chemists to evaluate the electron affinity of specific atoms in a compound. Recall from Chapter 6 that electronegativity indicates the relative ability of an atom to attract electrons in a chemical bond. Note that electronegativity values were assigned, whereas electron affinity values were measured.

Electronegativity The version of the periodic table of the elements shown in **Figure 8.20** lists electronegativity values. Note that fluorine has the greatest electronegativity value (3.98), while francium has the least (0.7). Because noble gases do not generally form compounds, individual electronegativity values for helium, neon, and argon are not listed. However, larger noble gases, such as xenon, sometimes bond with highly electronegative atoms, such as fluorine.

Figure 8.20 Electronegativity values are derived by comparing an atom's attraction for shared electrons to that of a fluorine's atom attraction for shared electrons. Note that the electronegativity values for the lanthanide and actinide series, which are not shown, range from 1.12 to 1.7.

1 H 2.20																	5 B 2.04	6 C 2.55	7 N 3.04	8 O 3.44	9 F 3.98	
3 Li 0.98	4 Be 1.57																	13 Al 1.61	14 Si 1.90	15 P 2.19	16 S 2.58	17 Cl 3.16
9 Na 0.93	12 Mg 1.31																	31 Ga 1.67	32 Ge 2.01	33 As 2.18	34 Se 2.55	35 Br 2.96
19 K 0.82	20 Ca 1.00	21 Sc 1.36	22 Ti 1.54	23 V 1.63	24 Cr 1.66	25 Mn 1.55	26 Fe 1.83	27 Co 1.88	28 Ni 1.91	29 Cu 1.90	30 Zn 1.65	31 Ga 1.67	32 Ge 2.01	33 As 2.18	34 Se 2.55	35 Br 2.96						
37 Rb 0.82	38 Sr 0.95	39 Y 1.22	40 Zr 1.33	41 Nb 1.6	42 Mo 2.16	43 Tc 2.19	44 Ru 2.2	45 Rh 2.28	46 Pd 2.20	47 Ag 1.93	48 Cd 1.69	49 In 1.78	50 Sn 1.96	51 Sb 2.05	52 Te 2.1	53 I 2.66						
55 Cs 0.79	56 Ba 0.89	57 La 1.10	72 Hf 1.3	73 Ta 1.5	74 W 1.7	75 Re 1.9	76 Os 2.2	77 Ir 2.2	78 Pt 2.2	79 Au 2.4	80 Hg 1.9	81 Tl 1.8	82 Pb 1.8	83 Bi 1.9	84 Po 2.0	85 At 2.2						
87 Fr 0.7	88 Ra 0.9	89 Ac 1.1																				
Metal Metalloid Nonmetal																						

Electronegativity Qnd Polarity Study Guide

Leo J. Malone, Theodore O. Dolter



Electronegativity Qnd Polarity Study Guide:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter A Self-study Guide to the Principles of Organic Chemistry Jiben Roy, 2013

A Self Study Guide to the Principles of Organic Chemistry Key Concepts Reaction Mechanisms and Practice Questions for the Beginner will help students new to organic chemistry grasp the key concepts of the subject quickly and easily as well as build a strong foundation for future study Starting with the definition of atom the author explains molecules electronic configuration bonding hydrocarbons polar reaction mechanisms stereochemistry reaction varieties organic spectroscopy aromaticity and aromatic reactions biomolecules organic polymers and a synthetic approach to organic compounds The over one hundred diagrams and charts contained in this volume will help students visualize the structures and bonds as they read the text and make the logic of organic chemistry clear and easily understood Each chapter ends with a list of frequently asked questions and answers followed by additional practice problems Answers are included in the Appendix **General,**

Organic, and Biochemistry Study Guide Marcia L. Gillette, Ira Blei, Wendy Gloffke, George Odian, 2006-01-27 This study guide provides reader friendly reinforcement of the concepts covered in the textbook Features include Chapter outlines Are you able to Worked text problems Fill ins Test yourself Concept maps Can also be used for Blei and Odian s Organic and Biochemistry **Organic Chemistry, 13e Student Study Guide and Solutions Manual** T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2022-05-03 Organic Chemistry Student Study Guide and Solutions Manual 13th Edition offers the full solutions for select exercises from the text **Organic Chemistry, 12e Binder Ready Version Study Guide & Student**

Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity

To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

AP Biology Study Guide Sundar Nathan,2009-11 Sundar Nathan received a Bachelor s degree in Electrical Engineering from Anna University Chennai India and a Masters degree in Biomedical Engineering from the University of Texas at Austin Working for over a year with a team of talented Phds MPhils and MScs from all over the world Sundar compiled this comprehensive study guide to help students prepare diligently understand the concepts and Crush the AP Bio Test

Chemistry, Student Study Guide James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Organic Chemistry Study Guide Robert J. Ouellette,J. David Rawn,2014-11-04 *Organic Chemistry Study Guide Key Concepts Problems and Solutions* features hundreds

of problems from the companion book Organic Chemistry and includes solutions for every problem Key concept summaries reinforce critical material from the primary book and enhance mastery of this complex subject Organic chemistry is a constantly evolving field that has great relevance for all scientists not just chemists For chemical engineers understanding the properties of organic molecules and how reactions occur is critically important to understanding the processes in an industrial plant For biologists and health professionals it is essential because nearly all of biochemistry springs from organic chemistry Additionally all scientists can benefit from improved critical thinking and problem solving skills that are developed from the study of organic chemistry Organic chemistry like any skill is best learned by doing It is difficult to learn by rote memorization and true understanding comes only from concentrated reading and working as many problems as possible In fact problem sets are the best way to ensure that concepts are not only well understood but can also be applied to real world problems in the work place Helps readers learn to categorize analyze and solve organic chemistry problems at all levels of difficulty Hundreds of fully worked practice problems all with solutions Key concept summaries for every chapter reinforces core content from the companion book *CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE* Kevin Smith,2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text

Study Guide to Accompany Basics for Chemistry Martha Mackin,2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students **Basic Concepts of Chemistry, Study Guide** Leo J. Malone,2000 The text s three main goals are to

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Table of Contents Electronegativity Qnd Polarity Study Guide

1. Understanding the eBook Electronegativity Qnd Polarity Study Guide
 - The Rise of Digital Reading Electronegativity Qnd Polarity Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Electronegativity Qnd Polarity Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electronegativity Qnd Polarity Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electronegativity Qnd Polarity Study Guide

- Personalized Recommendations
- Electronegativity Qnd Polarity Study Guide User Reviews and Ratings
- Electronegativity Qnd Polarity Study Guide and Bestseller Lists
- 5. Accessing Electronegativity Qnd Polarity Study Guide Free and Paid eBooks
 - Electronegativity Qnd Polarity Study Guide Public Domain eBooks
 - Electronegativity Qnd Polarity Study Guide eBook Subscription Services
 - Electronegativity Qnd Polarity Study Guide Budget-Friendly Options
- 6. Navigating Electronegativity Qnd Polarity Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Electronegativity Qnd Polarity Study Guide Compatibility with Devices
 - Electronegativity Qnd Polarity Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electronegativity Qnd Polarity Study Guide
 - Highlighting and Note-Taking Electronegativity Qnd Polarity Study Guide
 - Interactive Elements Electronegativity Qnd Polarity Study Guide
- 8. Staying Engaged with Electronegativity Qnd Polarity Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electronegativity Qnd Polarity Study Guide
- 9. Balancing eBooks and Physical Books Electronegativity Qnd Polarity Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electronegativity Qnd Polarity Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Electronegativity Qnd Polarity Study Guide
 - Setting Reading Goals Electronegativity Qnd Polarity Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electronegativity Qnd Polarity Study Guide

- Fact-Checking eBook Content of Electronegativity Qnd Polarity Study Guide
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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