

## Extension Questions

24. The hydrogen spectral lines in Model 2 are only the wavelengths of light that are in the visible range and therefore "seen" by the naked eye. However, many other wavelengths can be detected with special equipment.

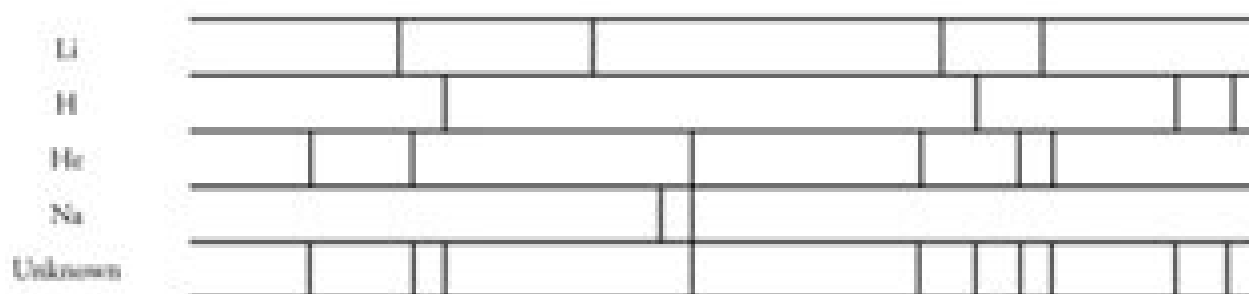
- a. Propose a hydrogen electron transition that involves light with a wavelength in the ultraviolet (UV) range (10–400 nm).

Wavelengths within the UV range would have more energy than violet wavelengths. This transition would be from  $n=6$  to  $n=2$  or maybe  $n=6$  to  $n=1$ .

- b. Propose a hydrogen electron transition that involves light with a wavelength in the infrared (IR) range (1000–10<sup>6</sup> nm).

Wavelengths within the IR range would have less energy than red wavelengths. This transition would be from  $n=3$  to  $n=2$  or maybe  $n=5$  to  $n=4$ .

25. Below are diagrams for the bright line spectra of four elements and the spectrum of a mixture of unknown gases.



- a. Which element(s) are not present in the Unknown?

# Electron Energy And Light Extension Questions Pogil Answers

**Guido Pizzella, C. D. Laughlin, Brian J.  
O'Brien**



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Photo-electricity Herbert Stanley Allen,1913      Photoelectricity - The Liberation of Electrons by Light H. Stanley Allen,2003-02      *Electron Energy Levels* ,1962      *Electron Energy Levels in Ordered and Disordered Structures* Marilyn E. Usher,1994      **Electron Energy Loss Spectra of Atoms and Molecules** ,1970      *Note on the Electron Energy Spectrum in the Inner Van Allen Belt* Guido Pizzella,C. D. Laughlin,Brian J. O'Brien,1962      **Electron energy levels in crystalline solids** David K. Adams,1952

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